

1. **What is the purpose of the following experiment?** (20%)
 The purpose of this experiment is to determine the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is exothermic, and the rate of reaction is measured by the volume of oxygen gas produced over time.

Temperature (°C)	Time (s)	Volume of O ₂ (cm ³)
20	10	10
20	20	20
20	30	30
20	40	40
20	50	50
20	60	60
20	70	70
20	80	80
20	90	90
20	100	100

Table 1: Data for the reaction of hydrogen peroxide and potassium iodide at 20°C.

Results and Discussion

- What is the rate of reaction at 20°C? (10%)
 The rate of reaction at 20°C is 1.0 cm³ s⁻¹.
- What is the rate of reaction at 30°C? (10%)
 The rate of reaction at 30°C is 2.0 cm³ s⁻¹.
- What is the rate of reaction at 40°C? (10%)
 The rate of reaction at 40°C is 4.0 cm³ s⁻¹.
- What is the rate of reaction at 50°C? (10%)
 The rate of reaction at 50°C is 8.0 cm³ s⁻¹.

2. **What is the effect of temperature on the rate of reaction?** (20%)
 The effect of temperature on the rate of reaction is that the rate of reaction increases as the temperature increases. This is because the molecules have more kinetic energy and are therefore more likely to collide with sufficient energy to overcome the activation energy barrier. The data shows that the rate of reaction increases by a factor of 8 when the temperature increases from 20°C to 50°C.

Temperature (°C)	Rate of Reaction (cm ³ s ⁻¹)
20	1.0
30	2.0
40	4.0
50	8.0

Pearson Modern Genetics Chapter Test Answer Key

Michael Seilmaier



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identify the direction that a river would flow on a topographic map · 2) compare two rivers/streams and determine ...

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