

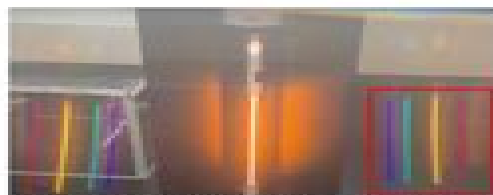
# Atomic Spectra Lab

## Chem 10

### PROCEDURE

#### Station 1: Emission Tubes

1. Put on a pair of yellow cardstock glasses from the box
2. Find the switch on the side of the lamp
3. Flip the switch to turn on the lamp
4. Write the color of the lamp in column A in the table below (ex: pale orange)
5. Look to right of the lamp to see the emission spectrum (inside the red box below)



6. Write the number of lines seen in the emission spectrum (ex: 4) in column B
7. Write the color of the lines seen in the emission spectrum (ex: dark blue, teal, yellow, hot pink) in column C
8. Turn off the lamp
9. Repeat steps 2-8 for the other two lamps
10. Return the yellow cardstock glasses to the cardboard box

Substance in the lamp	A: Color seen with eye	B: Number of lines observed using glasses (estimate)	C: Colors of lines in spectrum
He	Violet	3	Purple, blue, red
Ne	Red	20	Green, yellow, orange, red
Hg	Indigo	4	Yellow, green, blue, purple

#### Station 2: Burning Salts

1. Begin at the station where there is an empty petri dish (like the station shown below)

# Virtual Chem Lab 5 Atomic Emission Spectra

**Brian F. Woodfield, Matthew C.  
Asplund, Steven Haderlie**



## Virtual Chem Lab 5 Atomic Emission Spectra:

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