

## Writing Balanced Precipitation Reactions

In this problem, you will go into the virtual laboratory and perform a series of precipitation reactions using  $\text{Ag}^+$ ,  $\text{Pb}^{2+}$ , and  $\text{Sb}^{3+}$ . After observing the reactions, you will write the net ionic equations representing these reactions and then balance them.

1. Start *Virtual ChemLab*, select *Reactions and Stoichiometry*, and then select *Writing Balanced Precipitation Reactions* from the list of assignments. The lab will open in the Inorganic laboratory.
2. Drag a test tube from the box and place it on the metal test tube stand. You can then click on the bottle of  $\text{Ag}^+$  ion solution (it is  $\text{AgNO}_3$ ) on the shelf to add it to the test tube. Then drag the test tube to the blue test tube rack. Repeat for  $\text{Pb}^{2+}$  ( $\text{Pb}(\text{NO}_3)_2$ ) and  $\text{Sb}^{3+}$  ( $\text{Sb}(\text{OH})_3$ ).
3. Moving your mouse over each test tube in the test-tube rack will allow you to identify it on the TV monitor. You can also click on the label at the top of each test tube in order to label the test tubes. Place the test tube containing the  $\text{Ag}^+$  solution in the metal test tube stand. You will be able to observe what happens in the window at the top left. Click on the  $\text{Na}_2\text{CO}_3$  reagent bottle to add it to the test tube in the stand.

*What color is the precipitate? Write a correct balanced net ionic equation for the reaction.*

Brown

When you have finished writing the equation and making your observations, place the test tube containing the precipitate in the red disposal bucket.

4. Place the test tube containing the  $\text{Pb}^{2+}$  solution in the metal test tube stand. Click on the  $\text{NaCl}$  reagent bottle to add it to the test tube in the stand.

*What color is the precipitate? Write a correct balanced net ionic equation for this reaction.*

White

When you have finished writing the equation and making your observations, place the test tube containing the precipitate in the red disposal bucket.

5. Place the test tube containing the  $\text{Sb}^{3+}$  solution in the metal test tube stand. Note that the antimony is not soluble and you will have a double displacement reaction. Click on the  $\text{Na}_2\text{S}$  reagent bottle to add it to the test tube in the stand.

*What color is the precipitate? Write a correct balanced net ionic equation for the reaction.*

Orange

# Virtual Lab Writing Balanced Precipitation Reactions Answers

**Xiang Xie**



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## **Table of Contents Virtual Lab Writing Balanced Precipitation Reactions Answers**

1. Understanding the eBook Virtual Lab Writing Balanced Precipitation Reactions Answers
  - The Rise of Digital Reading Virtual Lab Writing Balanced Precipitation Reactions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Virtual Lab Writing Balanced Precipitation Reactions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Virtual Lab Writing Balanced Precipitation Reactions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Virtual Lab Writing Balanced Precipitation Reactions Answers
  - Personalized Recommendations
  - Virtual Lab Writing Balanced Precipitation Reactions Answers User Reviews and Ratings
  - Virtual Lab Writing Balanced Precipitation Reactions Answers and Bestseller Lists
5. Accessing Virtual Lab Writing Balanced Precipitation Reactions Answers Free and Paid eBooks
  - Virtual Lab Writing Balanced Precipitation Reactions Answers Public Domain eBooks
  - Virtual Lab Writing Balanced Precipitation Reactions Answers eBook Subscription Services
  - Virtual Lab Writing Balanced Precipitation Reactions Answers Budget-Friendly Options
6. Navigating Virtual Lab Writing Balanced Precipitation Reactions Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Virtual Lab Writing Balanced Precipitation Reactions Answers Compatibility with Devices
  - Virtual Lab Writing Balanced Precipitation Reactions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Highlighting and Note-Taking Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Interactive Elements Virtual Lab Writing Balanced Precipitation Reactions Answers
  8. Staying Engaged with Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Virtual Lab Writing Balanced Precipitation Reactions Answers
  9. Balancing eBooks and Physical Books Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Virtual Lab Writing Balanced Precipitation Reactions Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Setting Reading Goals Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Fact-Checking eBook Content of Virtual Lab Writing Balanced Precipitation Reactions Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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