

# Chemistry Review Packet

Name: \_\_\_\_\_

Date: \_\_\_\_\_

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 16.0 grams    B. 26.0 grams    C. 34.0 grams    D. 44.0 grams    E. 50.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A.  $\text{HNO}_3$     B.  $\text{Na}_2\text{CO}_3$     C.  $\text{HNO}_2$     D.  $\text{NaOH}$     E.  $\text{HCO}_2\text{H}$

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Silver    B. Baking soda    C. Natural gas    D. Neon gas    E. Water

4. What is the approximate percent composition, by mass, of hydrogen, in the compound ethanol,  $\text{C}_2\text{H}_6\text{O}$  (Molecular weight, in grams: C=12.01, H=1.008, O=16.00)

- A. 4.1%    B. 12.1%    C. 36.5%    D. 54.5%    E. 20.0%

5. The empirical formula of a certain hydrocarbon is  $\text{C}_2\text{H}_5$ , and it has a molecular weight equal to 60 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A.  $\text{C}_2\text{H}_5$     B.  $\text{C}_4\text{H}_{10}$     C.  $\text{C}_6\text{H}_{15}$     D.  $\text{C}_8\text{H}_{20}$     E.  $\text{C}_{10}\text{H}_{25}$

6. Which of the following aqueous solutions at 25°C has the highest normal boiling point?

- A. 0.20 M  $\text{CaCl}_2$     B. 0.20 M  $\text{NaCl}$     C. 0.5 M  $\text{KNO}_3$     D. 1.0 M  $\text{C}_6\text{H}_{12}\text{O}_6$     E. 0.20 M  $\text{MgCl}_2$

7. Consider the reaction:



How is this reaction best classified?

- A. Precipitation    B. Redox reaction    C. Single replacement    D. Combustion    E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 10.0 L of hydrogen gas reacts more than fully with excess nitrogen, how many liters of ammonia form?

- A. 1.00 L    B. 1.50 L    C. 10.0 L    D. 15.0 L    E. 20.0 L

# Reagents Chemistry Review Packet

**K Morrison**



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