Chapter 6 Equilibrium

Question 6.12:

A mixture of 1.57 mol of N_2 , 1.92 mol of H_2 and 8.13 mol of NH_3 is introduced into a 20 L reaction vessel at 500 K. At this temperature, the equilibrium constant, K_c for the reaction $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ is 1.7×10^2 . Is the reaction mixture at equilibrium? If not, what is the direction of the net reaction?

≤ Answer 6.12:

The given reaction is:

$$N_2(g) + 3H_2(g) \implies 2NH_3(g)$$

The given concentration of various species is

$$[N_2] = \frac{1.57}{20} \text{ mol } L^{-1} [H_2] = \frac{1.92}{20} \text{ mol } L^{-1}$$

 $\Rightarrow [NH_3] = \frac{8.13}{20} \text{ mol } L^{-1}$

Now, reaction quotient Q, is:

$$Qc = \frac{[NH_3]^2}{[N_2][N_2]^3}$$

$$= \frac{\left(\frac{8.13}{20}\right)^2}{\left(\frac{1.57}{20}\right)\left(\frac{1.92}{20}\right)^3} = 2.4 \times 10^3$$

Since, $Q_c \neq K_c$ the reaction mixture is not at equilibrium.

Again, $Q_c > K_c$ Hence, the reaction will proceed in the reverse direction.

Question 6.13:

The equilibrium constant expression for a gas reaction is,

$$K_c = \frac{[NH_3]^4 [O_2]^3}{[NO]^4 [H_2O]^6}$$

Write the balanced chemical equation corresponding to this expression.

Answer 6.13:

The balanced chemical equation corresponding to the given expression can be written as: $4NO(g) + 6H_1O(g) \implies 4NH_1(g) + 5O_1(g)$

World Of Chemistry Chapter 6

Joe Jackson

World Of Chemistry Chapter 6:

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