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**John H Haynes**



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mar 13 2023 as this suggests solubility is a relative term virtually nothing is completely insoluble in water as a rough guideline we can classify solutes as soluble insoluble or sparingly soluble on the basis of the number of moles of solute dissolved per liter of solution 0.10 mol/l soluble 0.01 mol/l insoluble 0.01 0.10 mol/l sparingly

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mar 11 2021 consider these slightly soluble salts pbs with  $K_{sp} = 8.4 \times 10^{-28}$  pbso<sub>4</sub> with  $K_{sp} = 1.8 \times 10^{-8}$  pbio<sub>3</sub> with  $K_{sp} = 2.6 \times 10^{-13}$  which is the most soluble calculate the solubility in moles/l for pbso<sub>4</sub> how many grams of pbso<sub>4</sub> dissolve in 1 l of solution



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using the graph you finished on worksheet 2 write an equation for the line depending on student ability you may provide formula for the class day 3 and 4 solubility lab overview in this lab the students will compare the dissolving rate of salt and sucrose in water at room temperature they will do this by collecting data on the time needed