

Unit 3: Acids and Bases

This unit activity will help you meet these chemistry goals:

- 1. **Science Inquiry.**—This unit explores the nature of acids, bases, acids and bases reacting, including their properties and behavior, and collecting information that will demonstrate your knowledge of acids and bases.
- 2. **Chemical Skills.**—This unit employs techniques for measuring and analyzing.

Introduction

Acids and bases are common in our daily lives. Acids are found in many foods, such as lemons, oranges, and vinegar. Bases are found in many household products, such as soap, bleach, and ammonia. Acids and bases are important in many chemical reactions. For example, acids and bases react to form salts and water. This reaction is called neutralization. Acids and bases are also important in many biological processes. For example, acids are found in the stomach and bases are found in the blood.

- Is the compound an acid or a base?
- Is the compound strong or weak?

In this activity, you will explore the nature of acids and bases and how they react. You will learn about the properties of acids and bases, how they are measured, and how they react with each other. You will also learn about the importance of acids and bases in many chemical reactions. You will learn about the properties of acids and bases, how they are measured, and how they react with each other. You will also learn about the importance of acids and bases in many chemical reactions.

Objectives and Analysis

Task 1: Acids and Bases in Real-World Applications

Many common acids and bases are found in everyday life. Acids are found in many foods, such as lemons, oranges, and vinegar. Bases are found in many household products, such as soap, bleach, and ammonia. Acids and bases are important in many chemical reactions. For example, acids and bases react to form salts and water. This reaction is called neutralization.

- 1. For the acids and bases listed below, determine whether they are strong or weak acids or bases. Write the chemical formula for each acid or base.

Strong acids and bases:

Hydrochloric acid (HCl) → hydrochloric acid + water → HCl + H₂O → H₃O⁺ + Cl⁻
Sulfuric acid (H₂SO₄) → sulfuric acid + water → H₂SO₄ + H₂O → H₃O⁺ + HSO₄⁻

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