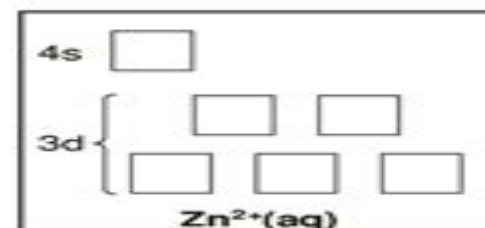
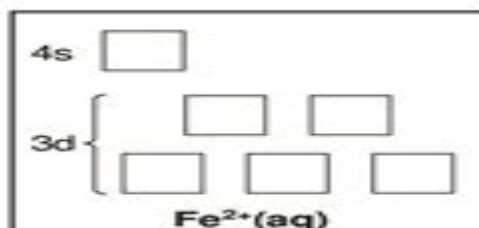
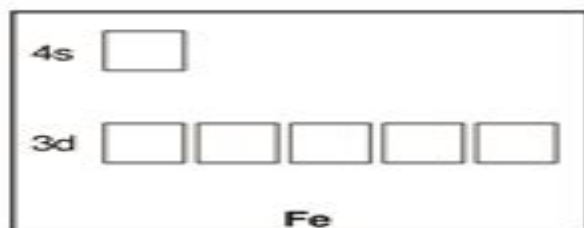


Section A

Answer **all** the questions in the spaces provided.

- 1 (a) (i) On the diagrams below, show the outer electron arrangements of the atoms and ions indicated. (Use the symbol $\uparrow\downarrow$ to represent a pair of electrons in an orbital.)



- (ii) Use the above diagrams to explain why Fe²⁺(aq) ions are coloured, whereas Zn²⁺(aq) ions are colourless.

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[4]

- (b) When concentrated HCl is added to a solution of Cu²⁺(aq) ions, the solution turns yellow.

- (i) State the formula of the species responsible for the yellow colour and name the *type of reaction* that has occurred.

.....

.....

- (ii) Ammonia can react as a base or as a ligand. Describe the colour changes that occur when NH₃(aq) is **gradually** added, with stirring, to the yellow solution, until the NH₃(aq) is in excess. Identify the **three** ions or compounds responsible for the new colours.

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[7]

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