

Name: _____

Neutralization Reactions Worksheet

Reading: Cycle 6 Chemistry II Lesson 4

Neutralization Reaction: Acid + Base \rightarrow Salt + Water

H⁺ from acid and OH⁻ from base combine to form H₂O. Leftovers form 'salt'.

Remember in salt, write metal first! KCl not ClK!

Example 1: $\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

Example 2: $__\text{HF} + __\text{Ca(OH)}_2 \rightarrow ______ + ______$

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)₂

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow ______ + ______$

Two H and two OH make two waters. Think of prom! Two boys and two girls make two couples.

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow \underline{2\text{ H}_2\text{O}} + ______$

What's left? 2 F atoms and 1 Ca atom form the 'salt'. Remember that Ca is the metal, so it goes first.

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow \underline{2\text{ H}_2\text{O}} + \underline{\text{CaF}_2}$

Safety check: is CaF₂ a good formula? P.T. tells you ions are Ca⁺² and F⁻, so total charge is 0 like it should be.

1. Complete & balance these neutralization reactions. Remember that H₂O = HOH!

$\text{H}_2\text{SO}_4 + \text{Sr(OH)}_2 \rightarrow __\text{H}_2\text{O} + \text{SrSO}_4$

$__\text{Ca(OH)}_2 + __\text{H}_3\text{PO}_4 \rightarrow __\text{H}_2\text{O} + __\text{Ca}_3(\text{PO}_4)_2$

$__\text{HBr} + \text{Ba(OH)}_2 \rightarrow __\text{H}_2\text{O} + ______$

$\text{Zn(OH)}_2 + __\text{HNO}_3 \rightarrow ______ + \text{Zn(NO}_3)_2$

$\text{Al(OH)}_3 + __\text{HCl} \rightarrow ______ + ______$

2. Complete and balance the following equations representing neutralization reactions:

a) $__\text{CsOH} + \text{H}_2\text{SO}_3 \rightarrow __\text{H}_2\text{O} + \text{Cs}_2\text{CO}_3$

b) $__\text{HF} + \text{Mg(OH)}_2 \rightarrow __\text{H}_2\text{O} + ______$

c) $__\text{HNO}_3 + \text{Al(OH)}_3 \rightarrow __\text{H}_2\text{O} + \text{Al(NO}_3)_3$

d) $\text{HCl} + \text{KOH} \rightarrow ______ + ______$

e) $\text{HBrO}_3 + \text{LiOH} \rightarrow \text{H}_2\text{O} + ______$

Neutralization Reactions Work Sheet 15

Crispin Collins



Neutralization Reactions Work Sheet 15:

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Table of Contents Neutralization Reactions Work Sheet 15

1. Understanding the eBook Neutralization Reactions Work Sheet 15
 - The Rise of Digital Reading Neutralization Reactions Work Sheet 15
 - Advantages of eBooks Over Traditional Books
2. Identifying Neutralization Reactions Work Sheet 15
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Neutralization Reactions Work Sheet 15
 - User-Friendly Interface
4. Exploring eBook Recommendations from Neutralization Reactions Work Sheet 15
 - Personalized Recommendations
 - Neutralization Reactions Work Sheet 15 User Reviews and Ratings
 - Neutralization Reactions Work Sheet 15 and Bestseller Lists

-
5. Accessing Neutralization Reactions Work Sheet 15 Free and Paid eBooks
 - Neutralization Reactions Work Sheet 15 Public Domain eBooks
 - Neutralization Reactions Work Sheet 15 eBook Subscription Services
 - Neutralization Reactions Work Sheet 15 Budget-Friendly Options
 6. Navigating Neutralization Reactions Work Sheet 15 eBook Formats
 - ePub, PDF, MOBI, and More
 - Neutralization Reactions Work Sheet 15 Compatibility with Devices
 - Neutralization Reactions Work Sheet 15 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Neutralization Reactions Work Sheet 15
 - Highlighting and Note-Taking Neutralization Reactions Work Sheet 15
 - Interactive Elements Neutralization Reactions Work Sheet 15
 8. Staying Engaged with Neutralization Reactions Work Sheet 15
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Neutralization Reactions Work Sheet 15
 9. Balancing eBooks and Physical Books Neutralization Reactions Work Sheet 15
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Neutralization Reactions Work Sheet 15
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Neutralization Reactions Work Sheet 15
 - Setting Reading Goals Neutralization Reactions Work Sheet 15
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Neutralization Reactions Work Sheet 15
 - Fact-Checking eBook Content of Neutralization Reactions Work Sheet 15
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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