

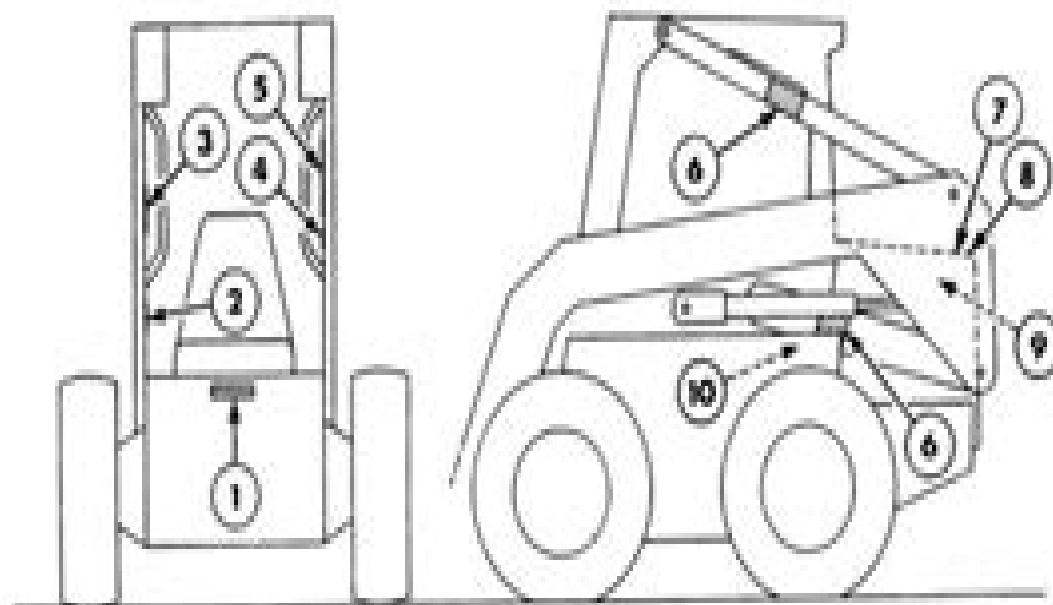
## SECTION 1 - GENERAL INFORMATION

## SAFETY DECALS

The following safety decals have been placed on your machine in the areas indicated. They are intended for the personal safety of you, and those working with you. Please take this manual, walk around your machine and note the content and

location of these warning signs. Review these decals with your machine operators.

Keep the decals legible. If they are not, obtain replacements from your New Holland dealer. The decal part numbers are listed with each decal.



**1. DANGER: DO NOT ALLOW PASSENGERS TO RIDE ON THE LOADER AT ANY TIME. DO NOT GET UNDER BOOM UNLESS SUPPORTED BY THE BOOM LOCK PINS.**

Part #00521005

# New Holland Skid Steer Lx885 Repair Manual

**JG Myers**



**New Holland Skid Steer Lx885 Repair Manual:**

*Rural Builder* ,1997      *Farmers and Consumers Market Bulletin* ,2008      *Service Manual, Skid Steer Loader L-555*  
*Sperry New Holland* *Sperry New Holland*,New Holland Machine Company,Ford Motor Company,1985      Repair Manual  
New Holland North America,1998      **New Holland repair manual: series TM 115-125-135-150-165 tractors - volume**  
**2** ,      *Ford 1320-1520 and 1720 Tractors Repair Manual* Ford New Holland Inc,198?      **Volvo 850 Service and Repair**  
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mol ko 2 and 0 mol h 2 o what is the limiting reactant

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$  x 1 mol  $\text{Al}$  x 2 mol  $\text{Al}_2\text{O}_3$  x 101.964 g  $\text{Al}_2\text{O}_3$  302.3 g  $\text{Al}_2\text{O}_3$  26.982 g  $\text{Al}$  4 mol  $\text{Al}$  1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield  $\text{yield actual} \times 100 = \frac{260.0 \text{ g}}{100.8601 \text{ theoretical}} \times 100 = 260.0 \text{ g}$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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