

- b) Calculate the density of NaCl, if edge length of NaCl unit cell is 564 pm. [Molar mass of NaCl = 58.5 g/mol]. (2)
[March 2013]
- 11) Unit cells can be broadly classified into 2 categories – primitive and centred unit cells.
- What is a unit cell? (1)
 - Name the three types of centred unit cells. (1 %)
 - The unit cell dimension of a particular crystal system is $a = b = c$ and $\alpha = \beta = \gamma = 90^\circ$. Identify the crystal system. (1)
 - Give one example for the above crystal system. (%) [SAY 2013]
- 12) a) Every substance has some magnetic properties associated with it. How will you account for the following magnetic properties?
- Paramagnetic property (1)
 - Ferromagnetic property (1)
- b) A compound is formed by two elements P and Q. Atoms of Q (as anions) make hcp lattice and those of the element P (as cations) occupy all the tetrahedral voids. What is the formula of the compound? (2) [March 2014]
- 13) a) Crystalline solids are 'anisotropic'. What is anisotropy? (1)
- b) Copper crystals have fcc unit cells.
- Compute the number of atoms per unit cell of copper crystals. (1)
 - Calculate the mass of a unit cell of copper crystals. (Atomic mass of Cu = 63.54u) (2) [SAY 2014]
- 14) Unit cells can be divided into two categories – primitive and centred unit cells.
- Differentiate between unit cell and crystal lattice. (2)
 - Calculate the number of atoms per unit cell in the following:
 - Body centred cubic unit cell (bcc)
 - Face centred cubic unit cell (fcc) (2) [March 2015]
- 15) a) Which of the following is not a characteristic of a crystalline solid?
- Definite heat of fusion
 - isotropic nature
 - A regular orderly arrangement of constituent particles
 - A true solid (1)
- b) Frenkel defect and Schottky defects are two stoichiometric defects found in crystalline solids.
- What are stoichiometric defects? (1)
 - Write any two differences between Schottky and Frenkel defects? (2) [SAY 2015]
- 16) a) Which of the following is a molecular solid?
- (a) Diamond (b) graphite (c) ice (d) quartz (1)
- (b) Unit cells can be classified into primitive and centered unit cells. Differentiate between primitive and centered unit cells. (1)
- (c) Presence of excess Na makes NaCl crystal coloured. Explain on the basis of crystal defects. (2) [March 2016]
- 17) A unit cell is a term related to crystal structure.
- What do you mean by unit cell? (1)
 - Name any two types of cubic unit cells. (1)
 - Calculate the number of atoms in each of the above mentioned cubic unit cells. (1)
 - Identify the substance which shows Frenkel defect:
 - NaCl
 - KCl
 - ZnS
 - AgBr (1) [SAY 2016]
- 18) a) Identify the non-stoichiometric defect
- Schottky defect
 - Frenkel defect
 - interstitial defect
 - Metal deficiency defect (1)
- b) What type of substance could make better permanent magnets – ferrimagnetic or ferromagnetic? Justify your answer. (2)

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