Step 2: Balance each kind of atom other that H and O
Fe²⁺ → Fe³⁺ + e⁻ Cr₂O₇²⁻ + 3e⁻ → 2Cr³⁺

Step 3: Balance oxygen atoms by using H2O

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
 $Cr_2O_7^{2-} + 3e^{-} \rightarrow 2Cr^{3+} + 7H_2O$

Step 4: Balance H atoms by using H+ ions

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
 $14H^{+} + Cr_{2}O_{7}^{2-} + 3e^{-} \rightarrow 2Cr^{3+} + 7H_{2}O_{7}^{2-}$

Step 5: Use electrons as needed to obtain a charge that is balanced

Fe²⁺
$$\rightarrow$$
 Fe³⁺ + e⁻ 14H⁺ + Cr₂O₇²⁻ +3e⁻ \rightarrow 2Cr³⁺ + 7H₂O

+2 +3 -1 +14 -2 -3 +6 0

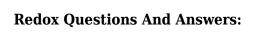
+2 +2 +9 +6

Add three electrons to the reactant side to balance the charges

$$Fe^{2+} \rightarrow Fe^{3+} + e^{-}$$
 which the the Garge $\Theta_7^{2-} + 6e^{-} \rightarrow 2Cr^{3+} + 7H_2O$

Redox Questions And Answers

Michael Rajnik



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