Vedantu

REDOX REACTION

Solved Examples

Example - 6 a

Here, two molecule of NaOH replaced 2H ion from the H₂SO₄. Therefore, each molecule of NaOH replaced only one H ion of acid, so v.f. = 1.

v. f. for acid is the number of OH replaced from the base by each molecule of acid

Solved Examples

Example - 7:

Here, one of molecule of H.SO, replaced one OH from NaOH. Therefore, valency factor for H.SO, is one

$$\therefore$$
 Eq. wt. of $H_sSO_4 = \frac{Mol. wt}{1}$

Salts :

(a) In non-reacting condition

Valency factor = Total number of positive charge or negative charge present in the compound.

Solved Examples

Example - 8:

$$\operatorname{Fe}_2(\operatorname{SO}_4)_3$$

Sed.

$$\frac{546}{2}$$

(b) In reacting condition

wit.

Solved Examples

Example - 9:

Sol. It is an acid base reaction, therefore valency factor for Na₂CO₃ is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of rodox change, v.f. = Total change in oxidation number per molecule.

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Arshad Iqbal

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