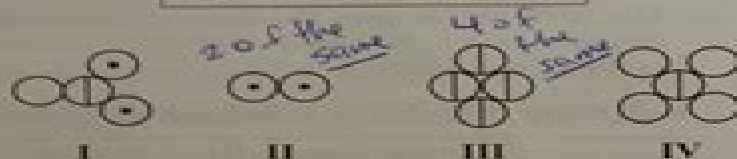
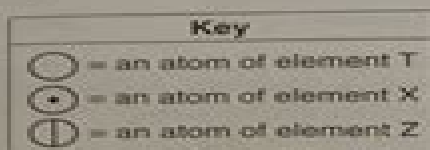


39 Given four particle models:



Which two models can be classified as elements?

- (1) I and II (3) II and III
(2) I and IV (4) II and IV

40 After being thoroughly stirred at $10.^{\circ}\text{C}$, which mixture is heterogeneous?

- (1) 25.0 g of KCl and 100. g of H_2O unsat.
(2) 25.0 g of KNO_3 and 100. g of H_2O supersat.
(3) 25.0 g of NaCl and 100. g of H_2O unsat.
(4) 25.0 g of NaNO_3 and 100. g of H_2O unsat.

41 Which two compounds are electrolytes?

- (1) KOH and CH_3COOH
(2) KOH and C_3H_{12}
(3) CH_3OH and CH_3COOH
(4) CH_3OH and C_3H_{12}

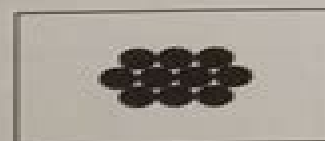
42 Which statement explains why a CO_2 molecule is nonpolar?

- (1) Carbon and oxygen are both nonmetals.
(2) Carbon and oxygen have different electronegativities.
(3) The molecule has a symmetrical distribution of charge.
(4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

- (1) $15.^{\circ}\text{C}$ to 298 K \uparrow (3) 305 K to $0.^{\circ}\text{C}$ \downarrow
(2) $37.^{\circ}\text{C}$ to 273 K \downarrow (4) 355 K to $25.^{\circ}\text{C}$ \downarrow

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? Solid

- (1) N_2 (g) (3) Mg (s)
(2) H_2 (g) (4) Kr (g)

45 Which type of equilibrium exists in a sealed flask containing $\text{Br}_2(\ell)$ and $\text{Br}_2(\text{g})$ at 298 K and 1.0 atm ?

- (1) static phase equilibrium
(2) static solution equilibrium
(3) dynamic phase equilibrium
(4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

- (1) $\text{KH}(\text{s})$, $\text{Cl}^-(\text{aq})$, and $\text{OH}^-(\text{aq})$
(2) $\text{K}(\text{s})$, $\text{Cl}_2(\text{g})$, and $\text{H}_2\text{O}(\ell)$
(3) $\text{KCl}(\text{aq})$ and $\text{H}_2\text{O}(\ell)$
(4) $\text{KOH}(\text{aq})$ and $\text{Cl}_2(\text{g})$

47 In a titration, 20.0 milliliters of a 0.150 M $\text{NaOH}(\text{aq})$ solution exactly neutralizes 24.0 milliliters of an $\text{HCl}(\text{aq})$ solution. What is the concentration of the $\text{HCl}(\text{aq})$ solution?

- (1) 0.125 M (3) 0.250 M
(2) 0.180 M (4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

- (1) $\frac{1}{2}$ (3) $\frac{1}{4}$
(2) $\frac{1}{3}$ (4) $\frac{1}{8}$

$$3 = \frac{87.3}{29.14}$$

$$\frac{1}{2} = \sqrt[3]{\frac{1}{8}}$$

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