

Name \_\_\_\_\_

Date \_\_\_\_\_

## Unit 10: Oxidation & Reduction – Introductory Review

### Oxidation States

1. What is the oxidation state of sulfur in  $\text{Na}_2\text{S}_2\text{O}_3$ ?

(1) +1

(2) +2

(3) +4

(4) +6

2. What is the oxidation number of chromium in  $\text{K}_2\text{Cr}_2\text{O}_7$ ?

(1) +12

(2) +6

(3) +3

(4) +8

3. What is the oxidation number of sulfur in  $\text{HSO}_3^-$ ?

10. Determine the change in oxidation number of oxygen during the reaction in test tube 1, shown below:



11. Common household bleach is an aqueous solution containing hypochlorite ions. A closed container of bleach is an equilibrium system represented by the equation below:



State the change in oxidation number for chlorine when the  $\text{Cl}_2(\text{g})$  changes to  $\text{Cl}^-(\text{aq})$  during the forward reaction.

# Redox And Electrochemistry Regents Short Answer Questions

**Wolfgang Guggemos**



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