

## REDOX REACTION

### Solved Examples

#### Example - 6 :



Base                  Acid

**Sol.** Valency factor of base = 1

Here, two molecule of NaOH replaced  $2\text{H}^+$  ion from the  $\text{H}_2\text{SO}_4$ . Therefore, each molecule of NaOH replaced only one  $\text{H}^+$  ion of acid, so v.f. = 1.

○ v. f. for acid is the number of  $\text{OH}^-$  replaced from the base by each molecule of acid

### Solved Examples

#### Example - 7 :



Base                  Acid

**Sol.** Valency factor of acid = 1

Here, one of molecule of  $\text{H}_2\text{SO}_4$  replaced one  $\text{OH}^-$  from NaOH. Therefore, valency factor for  $\text{H}_2\text{SO}_4$  is one

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

○ Salts :

(a) In non-reacting condition

○ Valency factor = Total number of positive charge or negative charge present in the compound.

### Solved Examples

#### Example - 8 :

		$\text{Na}_2\text{CO}_3$	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
<b>Sol.</b>	v.f.	2	$2 \times 3 = 6$	2
	Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

### Solved Examples

#### Example - 9 :



Base                  Acid

**Sol.** It is an acid base reaction, therefore valency factor for  $\text{Na}_2\text{CO}_3$  is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

# Redox Reactions Study Guide Answers

**Albert A Gayle**



## Redox Reactions Study Guide Answers:

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1. Understanding the eBook Redox Reactions Study Guide Answers
  - The Rise of Digital Reading Redox Reactions Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Redox Reactions Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Study Guide Answers
  - Personalized Recommendations
  - Redox Reactions Study Guide Answers User Reviews and Ratings
  - Redox Reactions Study Guide Answers and Bestseller Lists
5. Accessing Redox Reactions Study Guide Answers Free and Paid eBooks

- Redox Reactions Study Guide Answers Public Domain eBooks
- Redox Reactions Study Guide Answers eBook Subscription Services
- Redox Reactions Study Guide Answers Budget-Friendly Options
- 6. Navigating Redox Reactions Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Redox Reactions Study Guide Answers Compatibility with Devices
  - Redox Reactions Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Redox Reactions Study Guide Answers
  - Highlighting and Note-Taking Redox Reactions Study Guide Answers
  - Interactive Elements Redox Reactions Study Guide Answers
- 8. Staying Engaged with Redox Reactions Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Redox Reactions Study Guide Answers
- 9. Balancing eBooks and Physical Books Redox Reactions Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Redox Reactions Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Redox Reactions Study Guide Answers
  - Setting Reading Goals Redox Reactions Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Redox Reactions Study Guide Answers
  - Fact-Checking eBook Content of Redox Reactions Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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