EXPERIMENT 7: REDOX TITRATION

PURPOSE

The purpose of this experiment is to standardize the potassium permanganate solution by titration method.

INTRODUCTION

Oxidation-reduction processes electrons transfer one species to another. Potassium permanganate has been widely used as an oxidizing agent and requires no indicator unless very dilute solutions are used. The oxidizing action of KMnO₄ in the acidic medium can be represented by the following equation:

 $MnO_4 + 8H^* + 5e^- \longrightarrow Mn^{2*} + 4H_2O$

Solution containing MnO₄ ions are purple in colour and the solution containing Mn² ions colourless and hence permanganate solution is decolourised when added 10 a solution of a reducing agent.

CHEMICALS AND APPARATUS

Potassium permanganate solution	Burette	
Dried sodium oxalate	Conical flask	
IM sulphuric acid	Hot plate	
	Pipette	
	Thermometer	

PROCEDURE

- The burette was rinsed and filled with potassium permanganate solution, the air bubble was removed from the nozzle of the burette by draining some of the solutions...
- 0.3 g to 0.4 g of dried sodium oxalate was weighed accurately. The mass of sodium oxalate was noted and transferred to conical flask.
- 40 ml of 1.0 M sulphuric acid was added to the conical flask and carefully heated to 60°C.

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Peipei Pang

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