

4) Oxidation-Reduction Reactions

Oxidation - is a loss of electrons
 - is an increase in oxidation number

Reduction - is a gain of electrons
 - is a decrease in oxidation number

Oxidizing Agent - brings about the oxidation of another species
↑ gets reduced

Reducing Agent - brings about the reduction of another species
↑ causes something else to gain electrons by losing electrons itself (Is Oxidized)
↑ gets oxidized

★ OIL RIG (Oxidation Is Loss; Reduction Is Gain)
 LEO-GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal $\xrightarrow{\text{ox}}$ metal ion
 $2n \rightarrow 2n^{2+}$
2. Metal $\xrightarrow{\text{ox}}$ metallic ion $\xrightarrow{\text{ox}}$ metallic ion
 $\text{Cu} \rightarrow \text{Cu}^+ \rightarrow \text{Cu}^{2+}$
3. Halide ion $\xrightarrow{\text{ox}}$ halogen $\xrightarrow{\text{ox}}$ halite ion $\xrightarrow{\text{ox}}$ halate ion (similar for $\text{NO}_2^- \xrightarrow{\text{ox}} \text{NO}_3^-$)
 $\text{Cl}^- \rightarrow \text{Cl}_2 \rightarrow \text{ClO}^- \rightarrow \text{ClO}_2^- \rightarrow \text{ClO}_3^- \rightarrow \text{ClO}_4^-$

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ (acidified solution)
 $\text{MnO}_4^- \rightarrow \text{MnO}_2$ (basic solution)
 $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$ (acidified solution)

===Identify the substances being oxidized and reduced in each of the following.===



MEMORIZE!

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Table of Contents Redox And Electrochemistry Sheet Answers

1. Understanding the eBook Redox And Electrochemistry Sheet Answers
 - The Rise of Digital Reading Redox And Electrochemistry Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox And Electrochemistry Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox And Electrochemistry Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox And Electrochemistry Sheet Answers

- Personalized Recommendations
- Redox And Electrochemistry Sheet Answers User Reviews and Ratings
- Redox And Electrochemistry Sheet Answers and Bestseller Lists
- 5. Accessing Redox And Electrochemistry Sheet Answers Free and Paid eBooks
 - Redox And Electrochemistry Sheet Answers Public Domain eBooks
 - Redox And Electrochemistry Sheet Answers eBook Subscription Services
 - Redox And Electrochemistry Sheet Answers Budget-Friendly Options
- 6. Navigating Redox And Electrochemistry Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Redox And Electrochemistry Sheet Answers Compatibility with Devices
 - Redox And Electrochemistry Sheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox And Electrochemistry Sheet Answers
 - Highlighting and Note-Taking Redox And Electrochemistry Sheet Answers
 - Interactive Elements Redox And Electrochemistry Sheet Answers
- 8. Staying Engaged with Redox And Electrochemistry Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox And Electrochemistry Sheet Answers
- 9. Balancing eBooks and Physical Books Redox And Electrochemistry Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox And Electrochemistry Sheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Redox And Electrochemistry Sheet Answers
 - Setting Reading Goals Redox And Electrochemistry Sheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Redox And Electrochemistry Sheet Answers

- Fact-Checking eBook Content of Redox And Electrochemistry Sheet Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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