

## REDOX REACTION

### Solved Examples

#### Example - 6 :



Base                  Acid

**Sol.** Valency factor of base = 1

Here, two molecule of NaOH replaced  $2\text{H}^+$  ion from the  $\text{H}_2\text{SO}_4$ . Therefore, each molecule of NaOH replaced only one  $\text{H}^+$  ion of acid, so v.f. = 1.

○ v. f. for acid is the number of  $\text{OH}^-$  replaced from the base by each molecule of acid

### Solved Examples

#### Example - 7 :



Base                  Acid

**Sol.** Valency factor of acid = 1

Here, one of molecule of  $\text{H}_2\text{SO}_4$  replaced one  $\text{OH}^-$  from NaOH. Therefore, valency factor for  $\text{H}_2\text{SO}_4$  is one

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

○ Salts :

(a) In non-reacting condition

○ Valency factor = Total number of positive charge or negative charge present in the compound.

### Solved Examples

#### Example - 8 :

|             |         |                          |                              |   |
|-------------|---------|--------------------------|------------------------------|---|
|             |         | $\text{Na}_2\text{CO}_3$ | $\text{Fe}_2(\text{SO}_4)_3$ | $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$ |
| <b>Sol.</b> | v.f.    | 2                        | $2 \times 3 = 6$             | 2   |
|             | Eq. wt. | $\frac{M}{2}$            | $\frac{M}{6}$                | $\frac{M}{2}$                             |

(b) In reacting condition

### Solved Examples

#### Example - 9 :



Base                  Acid

**Sol.** It is an acid base reaction, therefore valency factor for  $\text{Na}_2\text{CO}_3$  is one while in non-reacting condition, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

# Redox Reactions Chemistry Study Guide Answers

**Sebastian Brünink**



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