Nia	me: Sammie Zishka
	Honors Chemistry Worksheet: Electrochemistry
1.)	From each pair, which is the better reducing agent? a.) Cd or (Fg) b.)(Ca) or Mg c.) Pb. (Zn) d.) Br., (Co)
	From each pair, which is the better excidizing agent? a.) (Ca) or Li* b.) From each pair, which is the better excidizing agent? b.) From each pair, which is the better excidizing agent? c.) (Cb) or Hg* d.) (Cu) or Cu2*
3	Consider the following reaction: $Mg_{00} + SnO_{Med} \longrightarrow MgSO_{g(e)} + SnO_{0}$
	a.) Which species is oxidized?
	h.) Which species is reduced?
	c.) What is the oxidizing reagent? ShO2.
	d.) What is the reducing reagent? N.Q.
4.	How can a person prevent corresion (exidation) of a steel tank that must be kept outdoors? Paint it. The paint is a barrier to prevent exidation.
	The ZING easily becomes a xiaized, which prevents the iron nails. Why does this work?
6.	Prom 0x1d)ZING. Describe the reactions that occurs at the smode and the cathode. Oxidation occurs at the anode. Reduction occurs at the Cathode. Cathode.
3.)	Use Table of Standard Reduction Potentials to determine the voltage for the following cells: a.) $Z_0(Z_0^{-\alpha}) = -3.2$
	b.) Mn(Mn2*1 Fe3* Fe 1.1 8 + -0.44 = 74.)
2	c.) ARAI" ASIAS" 66 + 80 = 2.40
Ю.	d) Niphi" Hgs" HglPt . 25 + 0.79 = 1.04 V
	e.) Ag and Ag* with Au and Au* - 8 + 1.5 = .7 V
	f.) Pb and Pb*2 with Cr*2 and Cr
	13 + .91 = .78V
63	Voltaic Cells 2.23
	a.) Draw a diagram of a Voltaic cell that corresponds to the following reaction:
	$M_{800} + Sn_{000}^{*2} \rightarrow Mg^{*00} + Sn_{00}$ $E' = 2.37 +$

Anode

 $Mg_{00} * Sn_{000} \longrightarrow Mg_{000} + Sn_{00}$ $2 \cdot 23 \text{ V}$ $N0^{+} \longrightarrow N0^{+}$ $S^{0} \longrightarrow S^{0}$

Cathode red.

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Table of Contents Reducction Potential Answer Key

- 1. Understanding the eBook Reducction Potential Answer Key
 - The Rise of Digital Reading Reducction Potential Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Reducction Potential Answer Key
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reducction Potential Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Reducction Potential Answer Key
 - Personalized Recommendations

- Reducction Potential Answer Key User Reviews and Ratings
- Reducction Potential Answer Key and Bestseller Lists
- 5. Accessing Reducction Potential Answer Key Free and Paid eBooks
 - Reducction Potential Answer Key Public Domain eBooks
 - Reducction Potential Answer Key eBook Subscription Services
 - Reducction Potential Answer Key Budget-Friendly Options
- 6. Navigating Reducction Potential Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Reducction Potential Answer Key Compatibility with Devices
 - Reducction Potential Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reducction Potential Answer Key
 - Highlighting and Note-Taking Reducction Potential Answer Key
 - Interactive Elements Reducction Potential Answer Key
- 8. Staying Engaged with Reducction Potential Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reducction Potential Answer Key
- 9. Balancing eBooks and Physical Books Reducction Potential Answer Key
 - ∘ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reducction Potential Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reducction Potential Answer Key
 - Setting Reading Goals Reducction Potential Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reducction Potential Answer Key
 - Fact-Checking eBook Content of Reducction Potential Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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