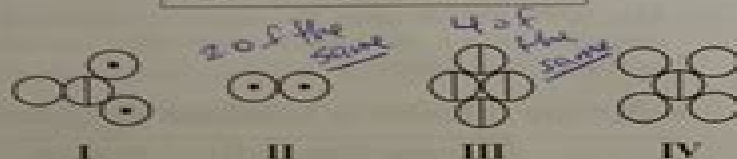
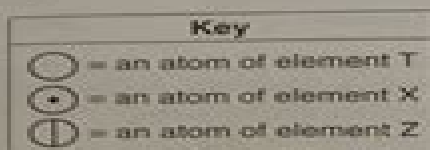


39 Given four particle models:



Which two models can be classified as elements?

- (1) I and II
 (2) I and IV
 (3) II and III
 (4) II and IV

40 After being thoroughly stirred at 10°C , which mixture is heterogeneous?

- (1) 25.0 g of KCl and 100. g of H_2O *unsat.*
 (2) 25.0 g of KNO_3 and 100. g of H_2O *supersat.*
 (3) 25.0 g of NaCl and 100. g of H_2O *unsat.*
 (4) 25.0 g of NaNO_3 and 100. g of H_2O *unsat.*

41 Which two compounds are electrolytes?

- (1) KOH and CH_3COOH
 (2) KOH and C_3H_{12}
 (3) CH_3OH and CH_3COOH
 (4) CH_3OH and C_3H_{12}

42 Which statement explains why a CO_2 molecule is nonpolar?

- (1) Carbon and oxygen are both nonmetals.
 (2) Carbon and oxygen have different electronegativities.
 (3) The molecule has a symmetrical distribution of charge.
 (4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

- (1) 15°C to 298 K \uparrow
 (2) 37°C to 273 K \downarrow
 (3) 305 K to 0°C \downarrow
 (4) 355 K to 25°C \downarrow

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? *Solid*

- (1) N_2 (g)
 (2) H_2 (g)
 (3) Mg (s)
 (4) Kr (g)

45 Which type of equilibrium exists in a sealed flask containing $\text{Br}_2(\ell)$ and $\text{Br}_2(\text{g})$ at 298 K and 1.0 atm ?

- (1) static phase equilibrium
 (2) static solution equilibrium
 (3) dynamic phase equilibrium
 (4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

- (1) $\text{KH}(\text{s})$, $\text{Cl}^-(\text{aq})$, and $\text{OH}^-(\text{aq})$
 (2) $\text{K}(\text{s})$, $\text{Cl}_2(\text{g})$, and $\text{H}_2\text{O}(\ell)$
 (3) $\text{KCl}(\text{aq})$ and $\text{H}_2\text{O}(\ell)$
 (4) $\text{KOH}(\text{aq})$ and $\text{Cl}_2(\text{g})$

47 In a titration, 20.0 milliliters of a 0.150 M $\text{NaOH}(\text{aq})$ solution exactly neutralizes 24.0 milliliters of an $\text{HCl}(\text{aq})$ solution. What is the concentration of the $\text{HCl}(\text{aq})$ solution?

- (1) 0.125 M
 (2) 0.180 M
 (3) 0.250 M
 (4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

- (1) $\frac{1}{2}$
 (2) $\frac{1}{3}$
 (3) $\frac{1}{4}$
 (4) $\frac{1}{8}$

Handwritten calculation:

$$3 = \frac{87.3}{29.14}$$

$$\frac{1}{2} = \frac{1}{8}$$

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Table of Contents Regents Chemistry Pearson Review Answers

1. Understanding the eBook Regents Chemistry Pearson Review Answers
 - The Rise of Digital Reading Regents Chemistry Pearson Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Regents Chemistry Pearson Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Regents Chemistry Pearson Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Regents Chemistry Pearson Review Answers
 - Personalized Recommendations
 - Regents Chemistry Pearson Review Answers User Reviews and Ratings

- Regents Chemistry Pearson Review Answers and Bestseller Lists
- 5. Accessing Regents Chemistry Pearson Review Answers Free and Paid eBooks
 - Regents Chemistry Pearson Review Answers Public Domain eBooks
 - Regents Chemistry Pearson Review Answers eBook Subscription Services
 - Regents Chemistry Pearson Review Answers Budget-Friendly Options
- 6. Navigating Regents Chemistry Pearson Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Regents Chemistry Pearson Review Answers Compatibility with Devices
 - Regents Chemistry Pearson Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Regents Chemistry Pearson Review Answers
 - Highlighting and Note-Taking Regents Chemistry Pearson Review Answers
 - Interactive Elements Regents Chemistry Pearson Review Answers
- 8. Staying Engaged with Regents Chemistry Pearson Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Regents Chemistry Pearson Review Answers
- 9. Balancing eBooks and Physical Books Regents Chemistry Pearson Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Regents Chemistry Pearson Review Answers
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Regents Chemistry Pearson Review Answers
 - Setting Reading Goals Regents Chemistry Pearson Review Answers
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Regents Chemistry Pearson Review Answers
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