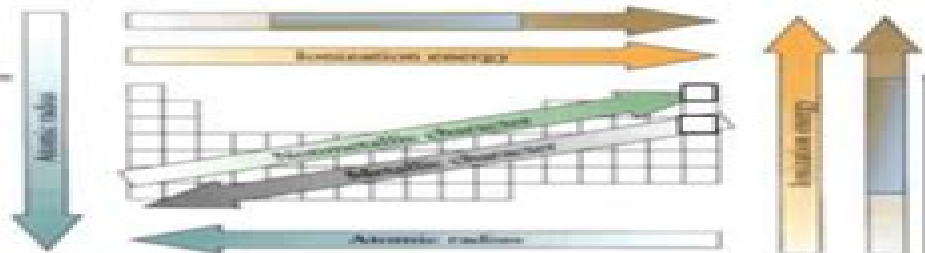


Chemistry Unit 5 – Periodic Law

Consider:

- How does location of an element on the periodic table relate to the chemical properties of that element?



Learning Targets: Students Will:

- ☐ SCL.CHEM.10.5.1 – know the family names that elements on the periodic table are placed into based on their electron configuration.
- ☐ SCL.CHEM.10.5.2 – recognize that chemical properties repeat on the periodic table due to similarity of valence electron configurations.
- ☐ SCL.CHEM.10.5.3 – know trends in atomic radius, ionization energy and electronegativity in the periodic table.
- ☐ SCL.CHEM.10.5.4 – predict relative reactivity of elements based on their location on the periodic table.
- ☐ SCL.CHEM.10.5.5 – predict ion charge of elements based on location on the periodic table.

KEY VOCABULARY:

- Periodic Law
- Nuclear Charge
- Atomic Radius
- Ionization Energy
- Electronegativity
- Ion
- Cation
- Anion
- Shielding Effect

How do we write an electron configuration?

m. ⁵⁸Ce

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 5d^1 4f^1$



KEY ACTIVITIES

- All in the Family
- Reactivity of Group 1 and 2 elements

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Raffaela Di Napoli



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