

- What is the oxidation number of chromium in the chromate ion, CrO_4^{2-} ?
A. -6
B. -2
C. +3
D. +6
- What is the oxidation number assigned to manganese in KMnO_4 ?
A. -7
B. -2
C. +3
D. +4
- Which changes occur when Fe^{2+} is reduced?
A. The Fe^{2+} gains electrons and its oxidation number increases.
B. The Fe^{2+} gains electrons and its oxidation number decreases.
C. The Fe^{2+} loses electrons and its oxidation number increases.
D. The Fe^{2+} loses electrons and its oxidation number decreases.
- Given the balanced ionic equation representing a reaction:



In this reaction, electrons are transferred from

- Al to Mg^{2+}
 - Al^{3+} to Mg
 - Mg to Al^{3+}
 - Mg^{2+} to Al
- Half-reactions can be written to represent all
A. double-replacement reactions
B. neutralization reactions
C. fission and fusion reactions
D. oxidation and reduction reactions
 - In an oxidation-reduction reaction, reduction is defined as the
A. loss of protons
B. gain of protons
C. loss of electrons
D. gain of electrons
 - In which reaction are electrons transferred from one reactant to another reactant?
A. $2\text{Ca}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{CaO}(\text{s})$
B. $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$
C. $\text{HCl}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
D. $\text{H}_3\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
 - Which balanced equation represents a redox reaction?
A. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$
B. $\text{BaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{BaCO}_3 + 2\text{KCl}$
C. $\text{CuO} + \text{CO} \rightarrow \text{Cu} + \text{CO}_2$
D. $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$
 - Which metal reacts spontaneously with a solution containing zinc ions?
A. magnesium
B. nickel
C. copper
D. silver

- Which reaction is an example of an oxidation-reduction reaction?
A. $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
B. $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$
C. $2\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$
D. $\text{Ba(OH)}_2 + 2\text{HCl} \rightarrow \text{BaCl}_2 + 2\text{H}_2\text{O}$

- Given the reaction:



Which species undergoes reduction?

- Al
 - Fe
 - Al^{3+}
 - Fe^{3+}
- According to Reference Table J, which of these metals will react most readily with 1.0 M HCl to produce $\text{H}_2(\text{g})$?
A. Ca
B. K
C. Mg
D. Zn
 - Given the balanced equation representing a redox reaction:



Which statement is true about this reaction?

- Each Al loses $2e^-$ and each Cu^{2+} gains $3e^-$.
 - Each Al loses $3e^-$ and each Cu^{2+} gains $2e^-$.
 - Each Al^{3+} gains $2e^-$ and each Cu loses $3e^-$.
 - Each Al^{3+} gains $3e^-$ and each Cu loses $2e^-$.
- Given the redox reaction:
$$\text{Cr}^{3+} + \text{Al} \rightarrow \text{Cr} + \text{Al}^{3+}$$

As the reaction takes place, there is a transfer of
A. electrons from Al to Cr^{3+}
B. electrons from Cr^{3+} to Al
C. protons from Al to Cr^{3+}
D. protons from Cr^{3+} to Al
 - Which half-reaction correctly represents reduction?
A. $\text{Ag} \rightarrow \text{Ag}^+ + e^-$
B. $\text{F}_2 \rightarrow 2\text{F}^- + 2e^-$
C. $\text{Au}^{3+} + 3e^- \rightarrow \text{Au}$
D. $\text{Fe}^{2+} + e^- \rightarrow \text{Fe}^{3+}$
 - Which equation shows conservation of both mass and charge?
A. $\text{Cl}_2 + \text{Br}^- \rightarrow \text{Cl}^- + \text{Br}_2$
B. $\text{Cu} + 2\text{Ag}^+ \rightarrow \text{Cu}^{2+} + \text{Ag}$
C. $\text{Zn} + \text{Cr}^{3+} \rightarrow \text{Zn}^{2+} + \text{Cr}$
D. $\text{Ni} + \text{Pb}^{2+} \rightarrow \text{Ni}^{2+} + \text{Pb}$
 - Which half-reaction can occur at the anode in a voltaic cell?
A. $\text{Ni}^{2+} + 2e^- \rightarrow \text{Ni}$
B. $\text{Sn} + 2e^- \rightarrow \text{Sn}^{2+}$
C. $\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$
D. $\text{Fe}^{3+} \rightarrow \text{Fe}^{2+} + e^-$

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