

Many neutralization reactions are exothermic. In **part 3** we will determine the heat of neutralization of hydrochloric acid with sodium hydroxide according to the reaction



This reaction takes place in the same calorimeter as above and we can summarize the heat flow the same way as we do in part 2: $-q_{\text{rxn}} = q_{\text{water}} + q_{\text{calorimeter}}$

Finally, we will measure the heat of fusion of ice in **part 4**. To accomplish that we add ice, initially at 0 °C, to the calorimeter with warm or room temperature water. The warmth of the water will be distributed to the

Part 3: Enthalpy of Neutralization Reaction

Combine about 25 mL of 3M NaOH with about 25 mL of 3 M HCl. Record volumes and temperatures.

	Trial 1	Trial 2	Trial 3
Volume of 3M HCl	24.0 mL	25.3 mL	25.8 mL
Volume of 3M NaOH	25.1 mL	27.1 mL	26.4 mL
T _i of solution	21.5 °C	22.6 °C	23.1 °C
T _f of solution	34.0 °C	33.2 °C	34.8 °C
ΔH_{rxn} in J/g			

Average ΔH_{rxn} in J/g: _____

Average ΔH_{rxn} in kJ/mol: _____

The standard enthalpy change for this reaction is -57.0 kJ/mol. Determine percent error in the experiment.

Percent error: _____

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