NEUTRALIZATION

The reaction of an acid with a base is called a "neutralization" reaction. There are two types that are emphasized here.

TYPE I.	Acid	+	Basic Hydroxide	0.00	H ₂ O	:=	Salt (ionic compound)
	EXAMPLES (balance these equations:)						
	HCI	+	NaOH	12	H_2O	+	NaCl
	HCl	+	Ca(OH) ₂	100	H_2O	*	CaCl ₂
	H ₂ SO ₄	+	кон	177	H ₂ O	+	K_2SO_4
	H ₂ SO ₄	+	Ba(OH) ₂	: **	H_2O	+	BaSO _a
	H_3PO_4	+	NaOH	1000	H_2O	+	Na_3PO_4
	H_3PO_4	+ -	Ca(OH) ₂	-	H_2O	-	$Ca_{3q}PO_4)_2$
TYPE 2.	Acid	+	Ammonia	0.400	Amm	oniu	nm Salt (ionic compound)
	EXAMPLES (balance these equations:)						
	HCI	+	NH_3	277	NH ₄ C	i	
	H ₂ SO ₄	+	NH ₃	100	(NH ₄)	so.	
	H,PO4	+	NH ₃	875	(NH ₄)	,PO	

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John M. Sheldon M.D. Professor of Internal Medicine Medical School E. Wendell Hewson, Ph.D. Professor of Meteorology College of Engineering

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Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

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