

- 2.74. According to Table 2.3, 1 gram of H_2O gives up 2260 J when it condenses. Since each gram of ice requires 333.5 J to melt, we get

$$2260 \text{ J} \times \frac{1 \text{ g}}{333.5 \text{ J}} = 6.78 \text{ grams of ice that can be melted.}$$

- 2.76. $_{\text{rxn}}H = \sum H(\text{prods}) - \sum H(\text{rts}) = (2 \text{ mol})(26.5 \text{ kJ/mol}) - 0 - 0 = 53.0 \text{ kJ}$

- 2.78. $_{\text{rxn}}H = 60_{\text{f}}H^\circ[\text{CO}_2(\text{g})] - [_{\text{f}}H^\circ[\text{C}_{60}(\text{s})] + 60_{\text{f}}H^\circ[\text{O}_2(\text{g})]] = 26367 \text{ kJ/mole}$

$$60(393.51 \text{ kJ/mol}) - 0 = 23610.6 \text{ kJ/mole}$$

$$_{\text{f}}H^\circ[\text{C}_{60}(\text{s})] = 2757 \text{ kJ/mole}$$

- 2.80. The reactions are:



This yields the overall reaction (you can verify that), and the overall $_{\text{rxn}}H$ is the sum of the values on the right: $_{\text{rxn}}H = 91.51 \text{ kJ}$.

- 2.82. The reaction is $2 \text{Al}(\text{s}) + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow \text{Al}_2\text{O}_3(\text{s}) + 2 \text{Fe}(\text{s})$. The $_{\text{rxn}}H$ is (using data from the appendix): $(2 \text{ mol})(0 \text{ kJ/mol}) + (1 \text{ mol})(-1675.7 \text{ kJ/mol}) - (1 \text{ mol})(-825.5 \text{ kJ/mol}) - (2 \text{ mol})(0 \text{ kJ/mol}) = -850.2 \text{ kJ}$.

- 2.84. In this case, since the combustion occurs in open atmosphere in (assumed) constant pressure, this time $q_p = H = -31,723 \text{ J}$. If the expansion of gases is done against a constant atmospheric pressure of 1 atm, to determine work we need to know the net volume of gas produced. From exercise 2.55, we found that there was a net change of -0.00491 mol of gas produced (ignoring the volume of the benzoic acid itself). Thus,

$$V = \frac{nRT}{P} = \frac{(0.0049 \text{ mol})(0.08205 \text{ L}\cdot\text{atm/mol}\cdot\text{K})(24.6 + 273.15)}{1 \text{ atm}} = 0.120 \text{ L} \quad \text{less gas volume}$$

Using this change in volume to calculate work:

$$w = P_{\text{ext}} \Delta V = (1 \text{ atm})(-0.120 \text{ L}) \left(\frac{101.32 \text{ J}}{1 \text{ L}\cdot\text{atm}} \right) = -12.1 \text{ J}$$

Finally, to calculate U :

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David W. Ball



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Table of Contents Physical Chemistry David Ball Solutions Manual

1. Understanding the eBook Physical Chemistry David Ball Solutions Manual
 - The Rise of Digital Reading Physical Chemistry David Ball Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry David Ball Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry David Ball Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry David Ball Solutions Manual

- Personalized Recommendations
 - Physical Chemistry David Ball Solutions Manual User Reviews and Ratings
 - Physical Chemistry David Ball Solutions Manual and Bestseller Lists
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 - Physical Chemistry David Ball Solutions Manual Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Chemistry David Ball Solutions Manual
 - Highlighting and Note-Taking Physical Chemistry David Ball Solutions Manual
 - Interactive Elements Physical Chemistry David Ball Solutions Manual
 8. Staying Engaged with Physical Chemistry David Ball Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry David Ball Solutions Manual
 9. Balancing eBooks and Physical Books Physical Chemistry David Ball Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry David Ball Solutions Manual
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Chemistry David Ball Solutions Manual
 - Setting Reading Goals Physical Chemistry David Ball Solutions Manual
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Chemistry David Ball Solutions Manual

- Fact-Checking eBook Content of Physical Chemistry David Ball Solutions Manual
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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