

Physical Science Chapter 7 Study Guide

- Molecules—groups of two or more atoms linked by chemical bonds to form distinct units
- Molecular mass—the sum of the mass of the atoms in a molecule
- Isomers—molecules with the same number and types of atoms but in a different arrangement
- Elements—substances composed of a single type of atom
- Compounds—substances composed of different types of atoms bonded together
- Mixtures—different kinds of atoms that are physically mixed but not chemically united
 - Solution—a mixture in which the ingredients are uniformly mixed
 - Solute—that which is dissolved in a substance
 - Solvent—the substance into which something is dissolved
 - Solubility—the ability of one substance to dissolve in another substance
 - Precipitate—solid particles that come out of a solution
 - Heterogeneous mixture—substances which are mixed but not completely or uniformly mixed
 - Colloid—a mixture that contains tiny suspended clumps of particles
- Chemical bonds—the attraction between atoms making up a molecule
 - Covalent—a bond which results from a sharing of electrons
 - Single covalent bonds means sharing one pair of electrons; double means two pairs, etc.
 - Lewis structure—a diagram that shows all of the atoms in a molecule and the valence electrons around each atom
 - Polar vs. nonpolar covalent bonds
 - Polar bond—the one atom has a stronger pull on the electrons (this is called its electronegativity) and therefore “hogs” the electrons most of the time, creating a negative electrical force around it and leaving the other atom with a positive electrical force
 - Nonpolar bond—the atoms in the molecule have a basically equal pull on the electrons; therefore the charge remains balanced
 - Ionic bond—attraction between oppositely charged atoms; occurs when one atom transfers electrons to another atom
 - Formula unit—The simplest ratio of atoms in a nonmolecular compound
 - Metallic bond—occurs between metal elements and involves the sharing of electrons around many nuclei (think Hutterite colony)
- Chemical formulas
 - Molecular formula—shows the exact number and kind of atoms in a molecule
 - Structural formula—shows the general arrangement of atoms in a molecule
 - Empirical formula—shows only the simplest ratio of atoms in molecule

Physical Science Chapter 7 Study Guide Answers

Paul A. Tipler, Gene Mosca



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Table of Contents Physical Science Chapter 7 Study Guide Answers

1. Understanding the eBook Physical Science Chapter 7 Study Guide Answers
 - The Rise of Digital Reading Physical Science Chapter 7 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 7 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 7 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 7 Study Guide Answers

- Personalized Recommendations
 - Physical Science Chapter 7 Study Guide Answers User Reviews and Ratings
 - Physical Science Chapter 7 Study Guide Answers and Bestseller Lists
5. Accessing Physical Science Chapter 7 Study Guide Answers Free and Paid eBooks
- Physical Science Chapter 7 Study Guide Answers Public Domain eBooks
 - Physical Science Chapter 7 Study Guide Answers eBook Subscription Services
 - Physical Science Chapter 7 Study Guide Answers Budget-Friendly Options
6. Navigating Physical Science Chapter 7 Study Guide Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Physical Science Chapter 7 Study Guide Answers Compatibility with Devices
 - Physical Science Chapter 7 Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Chapter 7 Study Guide Answers
 - Highlighting and Note-Taking Physical Science Chapter 7 Study Guide Answers
 - Interactive Elements Physical Science Chapter 7 Study Guide Answers
8. Staying Engaged with Physical Science Chapter 7 Study Guide Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 7 Study Guide Answers
9. Balancing eBooks and Physical Books Physical Science Chapter 7 Study Guide Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 7 Study Guide Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 7 Study Guide Answers
- Setting Reading Goals Physical Science Chapter 7 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 7 Study Guide Answers

- Fact-Checking eBook Content of Physical Science Chapter 7 Study Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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