

**QUESTION 3**

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a 0, 2 mol.dm<sup>-3</sup> ethanoic acid solution to 10 g of NaHCO<sub>3</sub> to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO<sub>2</sub> produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm<sup>3</sup> of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm<sup>3</sup> of a potassium hydroxide solution of concentration 0,25 mol.dm<sup>-3</sup> completely neutralises a dilute solution of sulphuric acid (H<sub>2</sub>SO<sub>4</sub>) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

**[25]**

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## **Table of Contents Physical Science Exam Papers N2**

1. Understanding the eBook Physical Science Exam Papers N2
  - The Rise of Digital Reading Physical Science Exam Papers N2
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Exam Papers N2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Exam Papers N2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Exam Papers N2

- Personalized Recommendations
- Physical Science Exam Papers N2 User Reviews and Ratings
- Physical Science Exam Papers N2 and Bestseller Lists
- 5. Accessing Physical Science Exam Papers N2 Free and Paid eBooks
  - Physical Science Exam Papers N2 Public Domain eBooks
  - Physical Science Exam Papers N2 eBook Subscription Services
  - Physical Science Exam Papers N2 Budget-Friendly Options
- 6. Navigating Physical Science Exam Papers N2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science Exam Papers N2 Compatibility with Devices
  - Physical Science Exam Papers N2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Exam Papers N2
  - Highlighting and Note-Taking Physical Science Exam Papers N2
  - Interactive Elements Physical Science Exam Papers N2
- 8. Staying Engaged with Physical Science Exam Papers N2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Exam Papers N2
- 9. Balancing eBooks and Physical Books Physical Science Exam Papers N2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Exam Papers N2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Exam Papers N2
  - Setting Reading Goals Physical Science Exam Papers N2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Exam Papers N2



- Fact-Checking eBook Content of Physical Science Exam Papers N2
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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