

Name: \_\_\_\_\_

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## Classifying Chemical Reactions Lab

**Educational Goals:** Students will understand...

- How to distinguish between a single displacement, double displacement, composition and decomposition reaction
- How to observe reactions and predict the product

**Student Objectives:** Students will...

- Perform a composition, decomposition, single displacement and double displacement reaction
- Make observations of chemical reactions and categorize them
- Write and balance chemical equations

### Pre-lab Questions

For each of the following experimental procedures and observations:

- a) Write a balanced chemical equation
- b) Name the product
- c) Identify the type of reaction

1. A student heats barium metal over a flame, and begins to react with oxygen gas in the air. A white, crystalline solid begins to form on the barium.

2. A chemist mixes lead (II) nitrate and potassium iodide solutions (both of which are clear liquids) and mixes them. A yellow precipitate appears in the final solution. The remaining liquid is an ionic solution. Solid precipitates containing lead are always yellow.

3. Magnesium metal is placed in sulfuric acid. The solution begins to bubble. The remaining liquid is an ionic solution.

# Physical Science If8767 Classifying Chemical Reactions

**Camilla Rothe**



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