Many neutralization reactions are exothermic. In part 3 we will determine the heat of neutralization of hydrochloric acid with sodium hydroxide according to the reaction

$$HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H2O(l)$$

This reaction takes place in the same calorimeter as above and we can summarize the heat flow the same way as we do in part 2: $-q_{exn} = q_{suter} + q_{calorimeter}$

Finally, we will measure the heat of fusion of ice in part 4. To accomplish that we add ice, initially at 0 °C, to the calorimeter with warm or room temperature water. The warmth of the water will be distributed to the

Part 3: Enthalpy of Neutralization Reaction

Combine about 25 mL of 3M NaOH with about 25 mL of 3 M HCl. Record volumes and temperatures.

Constant Constant Constant	Trial 1	Trial 2	Trial 3
Volume of 3M HCl	24.0 mL	25.3 mL	25.8 mL
Volume of 3M NaOH	25.1 mL	27.1 mL	26.4 mL
T _i of solution	21.5 °C	22.6 °C	23.1 °C
Tf of solution	34.0 °C	33.2 °C	34.8 °C
ΔH _{rxn} in J/g			1-000-000

Average ΔH _{rxn} in J/g:	
Average ΔH _{rsn} in kJ/mol:	
Average arras in Karmon.	

The standard enthalpy change for this reaction is -57.0 kJ/mol. Determine percent error in the experiment.

Percent	error:	
The second second		

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S K Jain

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