

# Neutralization-Titration II

## A. Molarity of an Unknown Acid

Example

Data Table

ONLY

	Sample 1		Sample 2		Sample 3 (if needed)	
	Acid*	Base	Acid*	Base	Acid*	Base
Final buret reading	X	20.0	X	37		
Initial buret reading	X	12.5	X	30		
Volume used	10 ml	7.5 ml		7.5 ml		

\*If a pipet is used to measure the volume of acid, record only in the space for volume used.

Molarity of base (NaOH) 0.2 m

CALCULATIONS: In the spaces below, show calculation setups for Sample 1 only. Show answers for both samples in the boxes.

$$(V)(M) = (V) \left( \frac{\text{moles}}{L} \right) = \text{moles}$$

1. Moles of base (NaOH) (if needed)

$$(7.5 \text{ mL}) \rightarrow \left( \frac{1L}{1000 \text{ mL}} \right) (0.2 \frac{\text{moles}}{L}) =$$

2. Moles of acid used to neutralize (react with) the above number of moles of base

moles HCl = moles NaOH when HCl solution is neutralized

3. Molarity of acid

$$M_{\text{HCl}} = \frac{\text{moles HCl}}{L_{\text{HCl}}} =$$

4. Average molarity of acid

5. Unknown acid number

Sample 1	Sample 2	Sample 3 (if needed)
<u>0.00198</u>	<u>0.00198</u>	
<u>0.00198 m</u>	<u>0.002 m</u>	
<u>0.1494 m</u>	<u>0.146 m</u>	

0.1472

# Neutralization And Titration Answers

**Princeton Review**



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