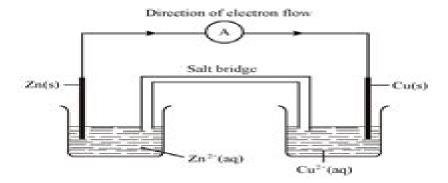
Redox Review - IB Short Questions

Question: 1



- (a) The apparatus shown above may be used to carry out a redox reaction.
 - (i) State the function of the salt bridge.
 - (ii) Write a half-equation for the exidation reaction.
 - (iii) The above reactions are carried out under standard conditions.

 State what the standard conditions are for the cell.
 - (iv) Using the Data Booklet, calculate the cell potential for the above cell. 2
 - (v) State and explain what happens to the concentration of the copper(II) ions when the cell is producing an electric current.
 - (vi) State two observations that could be made if the zine rod were placed in a solution of copper(II) ions.
- (b) The standard electrode potentials for three electrode systems are given below.

$$Ti^{3+}(aq) + c^{-} \rightarrow Ti^{2+}(aq)$$
 $E^{\bullet} = -0.37 \text{ V}$
 $Fc^{3+}(aq) + c^{-} \rightarrow Fc^{2+}(aq)$ $E^{\bullet} = +0.77 \text{ V}$
 $Cc^{4+}(aq) + c^{-} \rightarrow Cc^{3+}(aq)$ $E^{\bullet} = +1.45 \text{ V}$

- (i) Using the data above, deduce which species is the best reducing agent, giving a reason in terms of electrons for your answer.
- (ii) Write an equation, including state symbols, for the overall reaction with the greatest cell potential.

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