

MULTIPLE CHOICE QUESTIONS

Oxidation Number

Oxidation is a process of

- (A) loss of electrons
- (B) gain of electrons
- (C) increase in the negative valency
- (D) decrease in positive valency.

When zinc is added to CuSO_4 solution, copper is precipitated, it is so because of

- (A) Reduction of Zn
- (B) Reduction of Cu^{2+}
- (C) Hydrolysis of CuSO_4
- (D) Reduction of SO_4^{2-} .

The process in which oxidation number increases is

- (A) Reduction
- (B) Hydrolysis
- (C) Oxidation
- (D) Decomposition.

When zinc metal is added to dilute H_2SO_4 solution, hydrogen is evolved. In this zinc undergoes

- (A) Reduction
- (B) Oxidation
- (C) Dissolution
- (D) None of the above.

In which of the following, oxidation number of chlorine is +5 ?

- (A) Cl_2O_7
- (B) ClO_3^-
- (C) ClO^-
- (D) ClO_4^- .

Carbon has zero oxidation number in

- (A) CH_4
- (B) CH_3Cl
- (C) CCl_4
- (D) CH_2Cl_2 .

The oxidation number of carbon in $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ is

- (A) 0
- (B) +22
- (C) +6
- (D) -6.

In the following



- (A) Bromine is oxidised and carbonate is reduced
- (B) Bromine is reduced and carbonate is oxidised

- (C) Bromine is neither reduced nor oxidised
- (D) Bromine is reduced and oxidised.

9. The oxidation number of C in CH_4 , CH_3Cl , CH_2Cl_2 , CHCl_3 and CCl_4 is respectively

- (A) -4, -2, 0, +2, +4
- (B) +2, 4, 0, -2, -4
- (C) 4, 2, 0, -2, -4
- (D) 0, 2, -2, 4, 4.

10. Chlorine is in +3 oxidation state in

- (A) HCl
- (B) HClO_4
- (C) ICl
- (D) ClF_3 .

11. Oxidation number of Mn in K_2MnO_4 is

- (A) 2
- (B) 4
- (C) 6
- (D) 7.

12. A student states that burning of lime in air is an oxidation process. The reason he gives is that an oxide of the metal is produced on burning. Which one is correct ?

- (A) The statement and reason are true, the reason is correct explanation
- (B) The statement and reason are true, the explanation is not correct
- (C) The statement is true but the reason is false
- (D) The statement is false but the reason is true.

13. For the following reaction in the acidic solution



which of the following gives the true oxidation numbers of the manganese on each side of the equation ?

- (A) +7 to +6
- (B) +7 to +2
- (C) +4 to +2
- (D) -1 to +2.

14. In which of the following compounds, iron has the lowest oxidation number ?

- (A) $\text{Fe}(\text{CO})_5$
- (B) Fe_2O_3
- (C) $\text{K}_4[\text{Fe}(\text{CN})_6]$
- (D) $\text{FeSO}_4(\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$.

15. The oxidation number and covalency of sulphur in sulphur molecule (S_8) are

Answers

1. (A)
2. (B)
3. (C)
4. (B)
5. (B)
6. (D)
7. (A)
8. (D)
9. (A)
10. (D)
11. (C)
12. (D)
13. (B)
14. (A)

Redox Multiple Choice Questions And Answers

Angela Woodward Spangenberg



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Chemistry-I, **CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book **Teaching and Learning in the School Chemistry Laboratory** Avi Hofstein, Muhamad Hugerat, 2021-11-05 Research into the educational effectiveness of chemistry practical work has shown that the laboratory offers a unique mode of instruction assessment and evaluation Laboratory work is an integral and important part of the learning process used to encourage the development of high order thinking and learning alongside high order learning and thinking skills such as argumentation and metacognition Authored by renowned experts in the field of chemistry education this book provides a holistic approach to cover all issues related to learning and teaching in the chemistry laboratory With sections focused on developing the skill sets of teachers as well as approaches to supporting students in the laboratory the book offers a comprehensive look at vicarious instruction methods teacher and students roles and the blend with ICT simulations and other effective approaches to practical work The book concludes with a focus on retrospective issues followed up with a look to the future of laboratory learning A product of nearly fifty years of research this book will be useful for chemistry teachers curriculum developers researchers in chemistry education and professional development providers **AP Chemistry For Dummies** Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You'll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You'll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement

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web rotate the crankshaft pulley two complete turns clockwise and check the timing marks if the timing marks are correct tighten the lock bolt for the tensioner to 20 to 34 n m 200 to 350 kgfcm for fe to 37 to 52 n m 380 to 530 kgfcm for f2 engine check the movement on the tension side of the timing belt

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