

BALANCING REDOX REACTIONS

STEP BY STEP IN TWO WAYS...



BOTH START THE SAME WAY

1. ASSIGN OXIDATION NUMBERS TO EACH ATOM



NOTE: OXIDATION NUMBERS SUM UP TO ZERO EVERYWHERE

2. IDENTIFY THE OXIDATION AND REDUCTION HALF-REACTIONS

3. BALANCE EACH HALF-REACTION SEPARATELY

4. REVERSE YOUR EQUATION AND THE HALF-REACTIONS BALANCING ALL ATOMS INVOLVED IN IT

NOTE: REVERSE IT! + BECOMES - AND VICE-VERSA

OXIDATION: $\text{Fe}^{2+} \longrightarrow \text{Fe}^{3+}$

REDUCTION: $\text{SO}_4^{2-} \longrightarrow \text{SO}_2$

HALF-REACTION METHOD (in acidic conditions)

5. BALANCE OXIDATION

$\text{Fe}^{2+} \longrightarrow \text{Fe}^{3+} + 1\text{e}^-$

$\text{Fe}^{2+} + \text{SO}_4^{2-} \longrightarrow \text{SO}_2$

6. BALANCE REDUCTION

$10\text{H}^+ + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 2\text{H}_2\text{O}$

$\text{Fe}^{2+} + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 10\text{H}^+$

7. BALANCE CHARGES WITH ELECTRONS

NOTE: BALANCE CHARGES FIRST BY ADDING H⁺ AND THEN BALANCE THE ATOMS

$\text{Fe}^{2+} + 10\text{H}^+ + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 10\text{H}^+ + \text{Fe}^{3+}$

$-2 + 10 - 2 = 0$

$\text{Fe}^{2+} + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 10\text{H}^+ + \text{Fe}^{3+}$

$-2 = 0 + 10 - 2 = 8$

8. MAKE SURE THE NUMBER OF ELECTRONS IS THE SAME

NOTE: YOU CAN BALANCE EACH HALF-REACTION SEPARATELY UNDER ANY CONDITIONS

$\text{Fe}^{2+} + 10\text{H}^+ + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 10\text{H}^+ + \text{Fe}^{3+} \quad \times 2$

$\text{Fe}^{2+} + \text{SO}_4^{2-} \longrightarrow \text{SO}_2 + 10\text{H}^+ + \text{Fe}^{3+} \quad \times 2$

9. PUT BOTH REACTIONS BACK TOGETHER

$2\text{Fe}^{2+} + 2\text{SO}_4^{2-} + 2\text{H}^+ + 10\text{H}^+ + 2\text{Fe}^{3+} \longrightarrow$

$2\text{Fe}^{3+} + 10\text{H}^+ + 2\text{SO}_2 + 10\text{H}^+ + 2\text{Fe}^{3+}$

10. CHECK FOR ANYTHING THAT IS BOTH A PRODUCT AND REACTANT

NOTE: SOME REACTIONS CAN BE RUN WITH ANY OTHER IONS

$2\text{Fe}^{2+} + 2\text{SO}_4^{2-} + 2\text{H}^+ + 10\text{H}^+ + 2\text{Fe}^{3+} \longrightarrow$

$2\text{Fe}^{3+} + 10\text{H}^+ + 2\text{SO}_2 + 10\text{H}^+ + 2\text{Fe}^{3+}$



BUT WHAT ABOUT BASIC CONDITIONS?

WHEN WE SAY "ACIDIC" OR "BASIC" WE ARE USING TWO APPROXIMATE DEFINITIONS:

ACIDIC CONDITIONS HAVE A HIGH RATIO OF H⁺ TO OH⁻

IN BASIC WE USE THE SAME DEFINITION BUT ADD A REACTION REVERSING THE ACIDIC CONDITION

11. ADD THE NUMBER OF OH⁻ TO THE NUMBER OF H⁺ IN EACH SIDE

NOTE: ONLY ADD OH⁻ WHEN YOU DO THE SAME FOR BOTH HALF-REACTIONS

$2\text{Fe}^{2+} + 2\text{H}^+ + 2\text{SO}_4^{2-} + 2\text{Fe}^{3+} \longrightarrow 2\text{Fe}^{3+} + 2\text{SO}_2 + 2\text{H}^+ + 2\text{Fe}^{3+}$

$2\text{Fe}^{2+} + 2\text{SO}_4^{2-} + 2\text{Fe}^{3+} \longrightarrow 2\text{Fe}^{3+} + 2\text{SO}_2 + 2\text{Fe}^{3+} + 2\text{OH}^-$



Redox Reactions Teacher Guide And Answers

Liying Dong



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group II metals uses of halogens and their compounds The Halogenoalkanes Quiz Questions PDF e Book Chapter 14 interview questions and answers on Halogenoalkanes uses of halogenoalkanes elimination reactions nucleophilic substitution in halogenoalkanes and nucleophilic substitution reactions The Hydrocarbons Quiz Questions PDF e Book Chapter 15 interview questions and answers on Introduction to alkanes sources of alkanes addition reactions of alkenes alkane reaction alkenes and formulas The Introduction to Organic Chemistry Quiz Questions PDF e Book Chapter 16 interview questions and answers on Organic chemistry functional groups organic reactions naming organic compounds stereoisomerism structural isomerism and types of organic reactions The Ionic Equilibria Quiz Questions PDF e Book Chapter 17 interview questions and answers on Introduction to ionic equilibria buffer solutions equilibrium and solubility indicators and acid base titrations pH calculations and weak acids The Lattice Energy Quiz Questions PDF e Book Chapter 18 interview questions and answers on Introduction to lattice energy ion polarization lattice energy value atomization and electron affinity Born Haber cycle and enthalpy changes in solution The Moles and Equations Quiz Questions PDF e Book Chapter 19 interview questions and answers on Amount of substance atoms molecules mass chemical formula and equations gas volumes mole calculations relative atomic mass solutions and concentrations The Nitrogen and Sulfur Quiz Questions PDF e Book Chapter 20 interview questions and answers on Nitrogen gas nitrogen and its compounds nitrogen and gas properties ammonia ammonium compounds environmental problems caused by nitrogen compounds and nitrate fertilizers sulfur and oxides sulfuric acid and properties and uses of sulfuric acid The Organic and Nitrogen Compounds Quiz Questions PDF e Book Chapter 21 interview questions and answers on Amides in chemistry amines amino acids peptides and proteins The Periodicity Quiz Questions PDF e Book Chapter 22 interview questions and answers on Acidic oxides basic oxides aluminum oxide balancing equation period 3 chlorides balancing equations reactions with chlorine balancing equations reactions with oxygen bonding nature of period 3 oxides chemical properties of chlorine chemical properties of oxygen chemical properties periodicity chemistry periodic table chemistry oxides chlorides of period 3 elements electrical conductivity in period 3 oxides electronegativity of period 3 oxides ionic bonds molecular structures of period 3 oxides oxidation number of oxides oxidation numbers oxides and hydroxides of period 3 elements oxides of period 3 elements period III chlorides periodic table electronegativity physical properties periodicity reaction of sodium and magnesium with water and relative melting point of period 3 oxides The Polymerization Quiz Questions PDF e Book Chapter 23 interview questions and answers on Types of polymerization polyamides polyesters and polymer deductions The Rates of Reaction Quiz Questions PDF e Book Chapter 24 interview questions and answers on Catalysis collision theory effect of concentration reaction kinetics and temperature effect on reaction rate The Reaction Kinetics Quiz Questions PDF e Book Chapter 25 interview questions and answers on Reaction kinetics catalysts kinetics and reaction mechanism order of reaction rate constant k and rate of reaction The Redox Reactions and Electrolysis Quiz Questions PDF e Book Chapter 26 interview questions and answers on Redox reaction electrolysis

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Table of Contents Redox Reactions Teacher Guide And Answers

1. Understanding the eBook Redox Reactions Teacher Guide And Answers
 - The Rise of Digital Reading Redox Reactions Teacher Guide And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Teacher Guide And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Reactions Teacher Guide And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Teacher Guide And Answers
 - Personalized Recommendations
 - Redox Reactions Teacher Guide And Answers User Reviews and Ratings
 - Redox Reactions Teacher Guide And Answers and Bestseller Lists
5. Accessing Redox Reactions Teacher Guide And Answers Free and Paid eBooks
 - Redox Reactions Teacher Guide And Answers Public Domain eBooks
 - Redox Reactions Teacher Guide And Answers eBook Subscription Services
 - Redox Reactions Teacher Guide And Answers Budget-Friendly Options
6. Navigating Redox Reactions Teacher Guide And Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Redox Reactions Teacher Guide And Answers Compatibility with Devices
 - Redox Reactions Teacher Guide And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox Reactions Teacher Guide And Answers
 - Highlighting and Note-Taking Redox Reactions Teacher Guide And Answers
 - Interactive Elements Redox Reactions Teacher Guide And Answers
 8. Staying Engaged with Redox Reactions Teacher Guide And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox Reactions Teacher Guide And Answers
 9. Balancing eBooks and Physical Books Redox Reactions Teacher Guide And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox Reactions Teacher Guide And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Redox Reactions Teacher Guide And Answers
 - Setting Reading Goals Redox Reactions Teacher Guide And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Redox Reactions Teacher Guide And Answers
 - Fact-Checking eBook Content of Redox Reactions Teacher Guide And Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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