

Name _____

Date _____

Unit 10: Oxidation & Reduction – Introductory Review

Oxidation States

1. What is the oxidation state of sulfur in $\text{Na}_2\text{S}_2\text{O}_3$?

(1) +1

(2) +2

(3) +4

(4) +6

2. What is the oxidation number of chromium in $\text{K}_2\text{Cr}_2\text{O}_7$?

(1) +12

(2) +6

(3) +3

(4) +8

3. What is the oxidation number of sulfur in HSO_3^- ?

10. Determine the change in oxidation number of oxygen during the reaction in test tube 1, shown below:



11. Common household bleach is an aqueous solution containing hypochlorite ions. A closed container of bleach is an equilibrium system represented by the equation below:



State the change in oxidation number for chlorine when the $\text{Cl}_2(\text{g})$ changes to $\text{Cl}^-(\text{aq})$ during the forward reaction.

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