

- Which subatomic particles are located in the nucleus of a neon atom?
  - electrons and positrons
  - electrons and neutrons
  - protons and neutrons
  - protons and electrons
- Which statement best describes the nucleus of an aluminum atom?
  - It has a charge of +13 and is surrounded by a total of 10 electrons.
  - It has a charge of +13 and is surrounded by a total of 13 electrons.
  - It has a charge of -13 and is surrounded by a total of 10 electrons.
  - It has a charge of -13 and is surrounded by a total of 13 electrons.
- Which statement concerning elements is true?
  - Different elements must have different numbers of isotopes.
  - Different elements must have different numbers of neutrons.
  - All atoms of a given element must have the same mass number.
  - All atoms of a given element must have the same atomic number.
- What are the characteristics of a neutron?
  - It has no charge and no mass.
  - It has no charge and a mass of 1 amu.
  - It has a charge of +1 and no mass.
  - It has a charge of +1 and a mass of 1 amu.
- Which sequence represents a correct order of historical developments leading to the modern model of the atom?
  - the atom is a hard sphere → most of the atom is empty space → electrons exist in orbitals outside the nucleus
  - the atom is a hard sphere → electrons exist in orbitals outside the nucleus → most of the atom is empty space
  - most of the atom is empty space → electrons exist in orbitals outside the nucleus → the atom is a hard sphere
  - most of the atom is empty space → the atom is a hard sphere → electrons exist in orbitals outside the nucleus
- Which conclusion is based on the "gold foil experiment" and the resulting model of the atom?
  - An atom is mainly empty space, and the nucleus has a positive charge.
  - An atom is mainly empty space, and the nucleus has a negative charge.
  - An atom has hardly any empty space, and the nucleus has a positive charge.
  - An atom has hardly any empty space, and the nucleus has a negative charge.
- Subatomic particles can usually pass undeflected through an atom because the volume of an atom is composed of
  - an uncharged nucleus
  - largely empty space
  - neutrons
  - protons
- What is the net charge on an ion that has 9 protons, 11 neutrons, and 10 electrons?
  - 1+
  - 2+
  - 1-
  - 2-
- What is the total charge of the nucleus of a carbon atom?
  - 6
  - 0
  - +6
  - +12
- Which notation represents an atom of sodium with an atomic number of 11 and a mass number of 24?
  - ${}_{11}^{24}\text{Na}$
  - ${}_{24}^{11}\text{Na}$
  - ${}_{11}^{24}\text{Na}$
  - ${}_{24}^{11}\text{Na}$
- Compared to an atom of phosphorus-31, an atom of sulfur-32 contains
  - one less neutron
  - one less proton
  - one more neutron
  - one more proton
- Which subatomic particle has a negative charge?
  - proton
  - electron
  - neutron
  - positron
- Compared to a proton, an electron has
  - a greater quantity of charge and the same sign
  - a greater quantity of charge and the opposite sign
  - the same quantity of charge and the same sign
  - the same quantity of charge and the opposite sign
- How many electrons are in an  $\text{Fe}^{2+}$  ion.
  - 24
  - 26
  - 28
  - 56
- Which two particles make up most of the mass of a hydrogen-2 atom?
  - electron and neutron
  - electron and proton
  - proton and neutron
  - proton and positron
- Which two particles each have a mass approximately equal to one atomic mass unit?
  - electron and neutron
  - electron and positron
  - proton and electron
  - proton and neutron
- What is the mass number of an atom that has six protons, six electrons, and eight neutrons?
  - 6
  - 12
  - 14
  - 20
- What is the total number of neutrons in the nucleus of a neutral atom that has 19 electrons and a mass number of 39?
  - 19
  - 20
  - 39
  - 58
- In which list are the elements arranged in order of increasing atomic mass?
  - Cl, K, Ar
  - Fe, Co, Ni
  - Te, I, Xe
  - Ne, F, Na

# Regents Review Atomic Structure Answers

**KJ Lindholm-Leary**



## **Regents Review Atomic Structure Answers:**

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