

# 5<sup>th</sup> Grade Math Field Day Worksheets

## 5 Math Worksheets



Name \_\_\_\_\_



### Sports Themed Line Plots



Amount of miles ran at track practice

$X = 1$  Runner



How many total runners are represented in the line plot?

How many runners ran  $3/4$  mile or less?

How many more runners ran  $5/8$  mile than  $1/8$  mile?

What is the difference between the greatest distance ran and the least?

What is the combined distance that the runners ran at practice?



# Regional Math Field Day 5th Grade Practice

**Lei Shi**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe} + 2\text{Co} \rightarrow \text{Fe}^{2+} + 2\text{Co}^{2+}$  oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ag}^{+} + \text{Ni}$  oxidation reduction 25  $\text{Cu} + \text{Pb}^{2+} \rightarrow \text{Cu}^{2+} + \text{Pb}$  oxidation reduction 26  $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + 5\text{CO} \rightarrow 2\text{IO}_3^{-} + 5\text{CO}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation



state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^{+2}$  oxidized reducing agent  $\text{O}^0$  to  $\text{O}^{+2}$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^{+1}$  oxidized red ag  $\text{S}^0$  to  $\text{S}^{+2}$  reduced ox ag  $3\text{C} + 2\text{Br}_2 \rightarrow \text{C}_2\text{Br}_4$