

4) Oxidation-Reduction Reactions

Oxidation - is a loss of electrons
 - is an increase in oxidation number

Reduction - is a gain of electrons
 - is a decrease in oxidation number

Oxidizing Agent - brings about the oxidation of another species
↑ gets reduced

Reducing Agent - brings about the reduction of another species
↑ causes something else to gain electrons by losing electrons itself (Is Oxidized)
↑ gets oxidized

★ OIL RIG (Oxidation Is Loss; Reduction Is Gain)
 LEO-GER (Lose Electrons Oxidation ~ Gain Electrons Reduction)

Memorize common oxidizing pathways (reducing goes backward)

1. Metal $\xrightarrow{\text{ox}}$ metal ion
 $2n \rightarrow 2n^{2+}$
2. Metal $\xrightarrow{\text{ox}}$ metallic ion $\xrightarrow{\text{ox}}$ metallic ion
 $\text{Cu} \rightarrow \text{Cu}^+ \rightarrow \text{Cu}^{2+}$
3. Halide ion $\xrightarrow{\text{ox}}$ halogen $\xrightarrow{\text{ox}}$ halite ion $\xrightarrow{\text{ox}}$ halate ion (similar for $\text{NO}_2^- \xrightarrow{\text{ox}} \text{NO}_3^-$)
 $\text{Cl}^- \rightarrow \text{Cl}_2 \rightarrow \text{HOCl} \rightarrow \text{ClO}_2^- \rightarrow \text{ClO}_3^- \rightarrow \text{ClO}_4^-$

Memorize common oxidizing / reducing agents (see bright orange sheet)

Ex. $\text{MnO}_4^- \rightarrow \text{Mn}^{2+}$ (acidified solution)
 $\text{MnO}_4^- \rightarrow \text{MnO}_2$ (basic solution)
 $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+}$ (acidified solution)

===Identify the substances being oxidized and reduced in each of the following.===



MEMORIZE!

Redox And Electrochemistry Sheet Answers

John T. Moore, Richard H. Langley



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Ebook: Introductory Chemistry: An Atoms First Approach Burdge,2016-04-16 Ebook Introductory Chemistry An Atoms First Approach *Japanese Technical Periodical Index* ,1987

Chemistry-vol-I Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry DAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests Kaplan Test Prep,2016-10-04 Kaplan s DAT 2017 2018 Strategies

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