

## Chemistry Worksheets Class 11 on Chapter 8 Redox Reactions - Set 4

**Q-1:** Which of the following is not a redox reaction?

- a)  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- b)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- c) Photosynthesis
- d) Cell respiration

**Q-2:** Which of the following is not correct for an electrochemical cell?

- a) Convert electrical energy to chemical energy
- b) Anode is the negative terminal
- c) Anode is the positive terminal
- d) Cathode is the positive terminal

**Q-3:** Which reaction involves neither oxidation nor reduction?

- a)  $\text{CrO}_4^{2-} \rightarrow \text{Cr}_2\text{O}_7^{2-}$
- b)  $\text{Cr} \rightarrow \text{CrCl}_3$
- c)  $\text{KMnO}_4 \rightarrow \text{MnO}_2$
- d)  $\text{Fe} \rightarrow \text{Fe}_2\text{O}_3 \cdot x\text{H}_2\text{O}$

**Q-4:** Phosphorus has the oxidation state of +3 in

- a) Orthophosphoric acid
- b) Pyrophosphoric acid
- c) Metaphosphoric acid
- d) Phosphorous acid

**Q-5:** If a half cell reaction,  $\text{C} + \text{e}^- \rightarrow \text{C}^-$  has a large negative reduction potential, it follows that

- a) C is easily oxidised
- b) C is easily reduced
- c)  $\text{C}^-$  is readily oxidised
- d)  $\text{C}^-$  is readily reduced

**Q-6:** What is a redox couple?

**Q-7:** Identify the species undergoing oxidation and reduction in the following reactions.

- (i)  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- (ii)  $2\text{H}_2\text{S} + \text{O}_2 \rightarrow 2\text{S} + 2\text{H}_2\text{O}$
- (iii)  $\text{CH}_2=\text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_3-\text{CH}_3$

# Redox Reactions Worksheet 11 1 Answers

**Frank Schaffer Publications, Joan  
DiStasio**



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## **Table of Contents Redox Reactions Worksheet 11 1 Answers**

1. Understanding the eBook Redox Reactions Worksheet 11 1 Answers
  - The Rise of Digital Reading Redox Reactions Worksheet 11 1 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Redox Reactions Worksheet 11 1 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Redox Reactions Worksheet 11 1 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Reactions Worksheet 11 1 Answers
  - Personalized Recommendations
  - Redox Reactions Worksheet 11 1 Answers User Reviews and Ratings
  - Redox Reactions Worksheet 11 1 Answers and Bestseller Lists
5. Accessing Redox Reactions Worksheet 11 1 Answers Free and Paid eBooks
  - Redox Reactions Worksheet 11 1 Answers Public Domain eBooks
  - Redox Reactions Worksheet 11 1 Answers eBook Subscription Services
  - Redox Reactions Worksheet 11 1 Answers Budget-Friendly Options

6. Navigating Redox Reactions Worksheet 11 1 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Redox Reactions Worksheet 11 1 Answers Compatibility with Devices
  - Redox Reactions Worksheet 11 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Redox Reactions Worksheet 11 1 Answers
  - Highlighting and Note-Taking Redox Reactions Worksheet 11 1 Answers
  - Interactive Elements Redox Reactions Worksheet 11 1 Answers
8. Staying Engaged with Redox Reactions Worksheet 11 1 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Redox Reactions Worksheet 11 1 Answers
9. Balancing eBooks and Physical Books Redox Reactions Worksheet 11 1 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Redox Reactions Worksheet 11 1 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Redox Reactions Worksheet 11 1 Answers
  - Setting Reading Goals Redox Reactions Worksheet 11 1 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Redox Reactions Worksheet 11 1 Answers
  - Fact-Checking eBook Content of Redox Reactions Worksheet 11 1 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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