

8. Consider the following redox reaction: $2\text{MnO}_4^- + 3\text{ClO}_2 + \text{H}_2\text{O} \rightarrow 3\text{ClO}_2^- + 2\text{MnO}_2 + 2\text{OH}^-$

The reducing agent is:

- A) H_2O B) ClO_2^- C) MnO_2 D) MnO_4^-

9. A strong oxidizing agent has a:

- A) weak attraction for electrons B) strong attraction for electrons
C) weak ability to become reduced D) strong ability to become oxidized

10. Consider the following reaction: $3\text{As}_2\text{O}_3 + 4\text{NO}_3^- + 7\text{H}_2\text{O} \rightarrow 6\text{H}_2\text{AsO}_4^- + 4\text{HNO}_3$

The oxidizing agent is:

- A) H^+ B) H_2O C) NO_3^- D) As_2O_3

11. What is a typical characteristic of a strong oxidizing agent?

- A) It is readily oxidized B) It easily loses electrons.
C) It has a negative oxidation number. D) It has a positive reduction potential.

12. Two half-cells are connected via a salt bridge to make a voltaic cell. Which statement about this cell is correct?

- A) Oxidation occurs at the positive electrode (cathode).
B) It is also known as an electrolytic cell.
C) Ions flow through the salt bridge.
D) It requires a power supply to operate.

13. A solution containing an unknown cation reacts spontaneously with both zinc and copper. The unknown cation is:

- A) 1.0M H^+ B) 1.0M Ag^+ C) 1.0M Sr^{2+} D) 1.0M Ni^{2+}

14. As a metal corrodes:

- A) it gains electrons
B) it becomes reduced
C) it acts as a reducing agent
D) its oxidation number decreases

15. As a standard Zn / Ag electrochemical cell operates, in which direction do anions

move and how does the mass of the cathode change?

	Anion Direction	Mass of Cathode
A)	Towards Zn electrode	Increases
B)	Towards Ag electrode	Increases

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**Claudiu T. Supuran, William Alexander
Donald**



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Conversion of Water and CO₂ to Fuels using Solar Energy Oomman K. Varghese, Flavio L. Souza, 2024-02-05 Conversion of Water and CO₂ to Fuels using Solar Energy Comprehensive Resource for Understanding the Emerging Solar Technologies for Hydrogen Generation via Water Splitting and Carbon based Fuel Production via CO₂ Recycling Fossil fuel burning is the primary source of carbon in the atmosphere The realization that such burning can harm the life on our planet has led to a surge in research activities that focus on the development of alternative strategies for energy conversion Fuel generation using solar energy is one of the most promising approaches that has received widespread attention The fuels produced using sunlight are commonly referred to as solar fuels This book provides researchers interested in solar fuel generation a comprehensive understanding of the emerging solar technologies for hydrogen generation via water splitting and carbon based fuel production via CO₂ recycling The book presents the fundamental science technologies techno economic analysis and most importantly the materials that are being explored to

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