

## CHAPTER

## 08

## REDOX REACTION

- + The chemical reaction which contain oxidation as well as reduction half reaction is called redox reaction.
- + An atom or ion or molecule which undergoes reduction (gain of electrons) is called oxidant (while an atom or molecule which undergoes oxidation) (loss of electrons) is called reductant.
- + Oxygen shows positive oxidation states (+1 and +2) only in  $OF_2$  and  $O_2F_2$ .
- + Maximum oxidation state of an element never more than its group number.
- + a) Oxidation state of Mn in  $KMnO_4$  is +7.  
b) Oxidation state of Cr in  $K_2Cr_2O_7$  is +6.  
c) Oxidation state of Cr in  $CrO_3$  is +6.
- + a) Oxidation state of 'N' in  $N_3H$  is -1/3.  
b) Oxidation state of 'N' in  $NH_4NO_3$  ( $NH_4^+$  and  $NO_3^-$ ) are -3 and +5.
- + a) Oxidation state of 'Fe' in  $FeSO_4$  is +2.  
b) Oxidation state of 'Fe' in  $Fe_3O_4$  is +8/3.
- + In  $xMnO_4^- + yH_2O + zI^- \rightarrow 2MnO_2 + 8OH^- + 3I_2$   
 $x = 2, y = 4, z = 6$
- + In reaction of  $P_4 + xH_2O + yOH^- \rightarrow PH_3 + zH_2PO_2^-$   
 $x = 3, y = 3, z = 3$  (i.e.  $x = y = z$ )
- +  $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$  is an example for chemical combustion.
- + In all combustion reactions oxygen undergoes reduction (0 to -2).
- + In decomposition reaction chemical compound split into two or more simpler substances.  
 $[2H_2O \rightarrow 2H_2 + O_2]$
- + In decomposition of  $MgCO_3$ , no changes in oxidation numbers. So it is not a redox reaction  
 $[MgCO_3 \rightarrow MgO + CO_2]$
- +  $Zn_{(s)} + CuSO_{4(aq)} \rightarrow ZnSO_{4(aq)} + Cu_{(s)}$  is a displacement reaction.
- + In disproportionation reaction the same element undergoes both oxidation and reduction simultaneously.
- + When  $Br_2$  and  $I_2$  dissolve in  $CCl_4$  solvent to give reddish brown and purple colours respectively. So, they are detected by extracting into  $CCl_4$ . This is also called "Layer Test".
- + Comproportion is a reverse of disproportionation reaction

# Redox Reactions Chapter Study

**Adam Liwo**



## Redox Reactions Chapter Study:

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*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Computational Methods to Study the Structure and Dynamics of Biomolecules and Biomolecular Processes Adam Liwo, 2018-12-19 This book provides a comprehensive overview of modern computer based techniques for analyzing the structure properties and dynamics of biomolecules and biomolecular processes It is organized in four main parts the first one deals with methodology of molecular simulations the second one with applications of molecular simulations the third one introduces bioinformatics methods and the use of experimental information in molecular simulations the last part reports on selected applications of molecular quantum mechanics This second edition has been thoroughly revised and updated to include the latest progresses made in

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methods that assess the eff *Fundamentals of Environmental Sampling and Analysis* Chunlong Zhang, 2024-02-29

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