

## Electrochemistry Regents Questions

### Chemistry Regents June 2008

26 Which half-reaction equation represents the reduction of a potassium ion?



A flashlight can be powered by a rechargeable nickel-cadmium battery. In the battery, the anode is  $\text{Cd}(s)$  and the cathode is  $\text{NiO}_2(s)$ . The unbalanced equation below represents the reaction that occurs as the battery produces electricity. When a nickel-cadmium battery is recharged, the reverse reaction occurs.



75 Determine the change in oxidation number for the element that makes up the anode in the reaction that produces electricity. [1]

From \_\_\_\_\_ to \_\_\_\_\_

76 Explain why Cd would be above Ni if placed on Table J. [1]

51. Base your answer on the information below.

The unbalanced equation below represents the decomposition of potassium chlorate.



a Balance the equation in the space provided or on a separate piece of paper, using the smallest whole-number coefficients. [1]

b Determine the oxidation number of chlorine in the reactant. [1]



# Regents Questions On Electrochemistry

**G Thomas**



## **Regents Questions On Electrochemistry:**

## Unveiling the Magic of Words: A Overview of "**Regents Questions On Electrochemistry**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Regents Questions On Electrochemistry**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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