

Thermochemistry Chemistry Questions with Solutions

Q-1: Given that standard molar enthalpies of formation of NO(g) and $\text{NO}_2\text{(g)}$ are respectively 90.3 kJ/mol and 33.2 kJ/mol. Calculate the enthalpy change for the reaction $2\text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{NO}_2\text{(g)}$.

Answer: $\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$

$$\Delta H = 2(33.2) - [2(90.3) + 0]$$

$$\Delta H = -114.2 \text{ kJ}$$

Note: Enthalpy of formation of free elements like O_2 is always 0.

Q-2: The enthalpy of formation of AgCl is obtained from the enthalpy change from which one of the following process?

- a) $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl(s)}$
- b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$
- c) $\text{AgCl(s)} \rightarrow \text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g})$
- d) $\text{Ag(s)} + \text{AuCl} \rightarrow \text{Au(s)} + \text{AgCl(s)}$

Answer: b) $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Explanation: The enthalpy change accompanying the formation of one mole of a pure substance in its standard state from the constituents elements in their standard state is called enthalpy of formation. For Ag, the standard state is solid and for chlorine, it is gas. Thus, the process that gives the enthalpy of formation of AgCl is: $\text{Ag(s)} + \frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{AgCl(s)}$

Q-3: ΔH and ΔE for the reaction, $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \rightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(l)}$ at constant temperature are related as

- a) $\Delta H = \Delta E$
- b) $\Delta H = \Delta E + RT$
- c) $\Delta H = \Delta E + 3RT$
- d) $\Delta H = \Delta E - 3RT$

Answer: d) $\Delta H = \Delta E - 3RT$

Explanation: For any chemical reaction,

$$\Delta H = \Delta E + \Delta_{\text{ng}}RT$$

Where, Δ_{ng} = total number of moles of gaseous product- total number of moles of gaseous reactants.

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