

Davis P2

# Periodic Table of the Elements

Handwritten notes:

- Valence Electrons
- groups (columns)
- periods (rows)
- Mass # = protons + neutrons
- metals
- non-metals
- metalloids

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H	2 He																
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
55 Cs	56 Ba	57-71 Lanthanides	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn
87 Fr	88 Ra		89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

<ul style="list-style-type: none"> <li>1: alkali metals</li> <li>2: alkaline earth metals</li> <li>3: transition metals</li> <li>4: noble gases</li> </ul>	<p>by color/number</p> <ul style="list-style-type: none"> <li>1: metalloids</li> <li>2: non-metals</li> <li>3: noble gases</li> <li>4: transition metals</li> </ul>	<ul style="list-style-type: none"> <li>5: halogens</li> <li>6: noble gases</li> <li>7: non-metals</li> </ul>	<p>Other box</p> <ul style="list-style-type: none"> <li>1: solid</li> <li>2: liquid</li> <li>3: gas</li> </ul>
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# Reinforcement And Study Guide The Periodic Table

**Ralph M. Feather, Susan Leach  
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