

Naming Compounds Containing Polyatomic Ions

Use the following list of polyatomic ions to name the compounds given.

OH
NO₃
C₂H₃O₂

hydroxide
nitrate
acetate

ClO₃
HCO₃
SO₄
CO₃

chlorate
bicarbonate
sulfate
carbonate

SO₃
PO₄
NH₄

sulfite
phosphate
ammonium

- | | |
|--|---|
| 1. CuSO ₄ _____ | 26. NaOH _____ |
| 2. KClO ₃ _____ | 27. NaNO ₃ _____ |
| 3. NH ₄ OH _____ | 28. Ba(OH) ₂ _____ |
| 4. K ₂ CO ₃ _____ | 29. K ₂ CO ₃ _____ |
| 5. Na ₂ SO ₄ _____ | 30. LiOH _____ |
| 6. KC ₂ H ₃ O ₂ _____ | 31. KHCO ₃ _____ |
| 7. NH ₄ Cl _____ | 32. NH ₄ F _____ |
| 8. Mg(OH) ₂ _____ | 33. KClO ₃ _____ |
| 9. Zn(NO ₃) ₂ _____ | 34. NaC ₂ H ₃ O ₂ _____ |
| 10. MgSO ₄ _____ | 35. CaClO ₃ _____ |
| 11. CaCO ₃ _____ | 36. SrNO ₃ _____ |
| 12. KOH _____ | 37. Be(OH) ₂ _____ |
| 13. Ca(OH) ₂ _____ | 38. Li ₂ SO ₄ _____ |
| 14. K ₂ CO ₃ _____ | 39. Al ₂ (SO ₄) ₃ _____ |
| 15. AlPO ₄ _____ | 40. BaSO ₃ _____ |
| 16. Na ₂ PO ₄ _____ | 41. NH ₄ I _____ |
| 17. MgSO ₄ _____ | 42. NH ₄ HCO ₃ _____ |
| 18. AgNO ₃ _____ | 43. Cu ₂ (PO ₄) ₃ _____ |
| 19. Na ₂ CO ₃ _____ | 44. Ni(OH) ₂ _____ |
| 20. NaHCO ₃ _____ | 45. ZnSO ₄ _____ |
| 21. Ca(NO ₃) ₂ _____ | 46. MnCO ₃ _____ |
| 22. BaSO ₄ _____ | 47. H ₂ SO ₄ _____ |
| 23. BPO ₃ _____ | 48. Au ₂ SO ₄ _____ |
| 24. LiNO ₃ _____ | 49. NH ₄ NO ₃ _____ |
| 25. Al(NO ₃) ₃ _____ | 50. LiClO ₃ _____ |

Polyatomic Ions Naming Answer

Heather Hattori, Richard H. Langley



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Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make

available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *General, Organic, and Biological Chemistry* Michael P. Garoutte, 2014-02-24

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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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