

- (b) A solution is prepared by the addition of 10.0 mmol of  $\text{CH}_3\text{COOH}$  to 10.0 mmol of  $\text{CH}_3\text{COONa}$  and enough water to make 1.00 L of solution.

- (c) The solution reacts with a proper amount of  $\text{HCl}$  to produce a pH of 4.0.

How many moles of  $\text{HCl}$  are added to the solution that will produce a pH of 4.0?

In your answer, you should provide details of the general procedure that would be used in including appropriate equations and necessary calculations.

1. Use Henderson-Hasselbalch equation

$$\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$$

$$4.0 = 4.75 + \log \frac{[\text{A}^-]}{[\text{HA}]} \Rightarrow \frac{[\text{A}^-]}{[\text{HA}]} = 10^{4.0 - 4.75} = 10^{-0.75} = 0.1778$$

$$\frac{[\text{A}^-]}{[\text{HA}]} = \frac{0.0100 - x}{0.0100 + x} = 0.1778 \Rightarrow x = 0.0041 \text{ mol}$$

$$\text{Moles of HCl added} = 0.0041 \text{ mol} \times 1.00 \text{ L} = 0.0041 \text{ mol}$$

Reaction:  $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$   
Initial: 0.0100 mol, 0.0100 mol, 0.0000 mol, 0.0000 mol  
Change: -x, -x, +x, +x  
Equilibrium: 0.0059 mol, 0.0059 mol, 0.0041 mol, 0.0041 mol  
pH = 4.0

- (b) A buffer is prepared by the addition of 10.0 mmol of  $\text{CH}_3\text{COOH}$  to 10.0 mmol of  $\text{CH}_3\text{COONa}$  and enough water to make 1.00 L of solution.

How many moles of  $\text{HCl}$  are added to the solution that will produce a pH of 4.0?



- (c) The solution reacts with a proper amount of  $\text{HCl}$  to produce a pH of 4.0.

How many moles of  $\text{HCl}$  are added to the solution that will produce a pH of 4.0?

- (d) The solution reacts with a proper amount of  $\text{HCl}$  to produce a pH of 4.0.

$$\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]} \Rightarrow \frac{[\text{A}^-]}{[\text{HA}]} = 10^{4.0 - 4.75} = 0.1778$$

$$\frac{[\text{A}^-]}{[\text{HA}]} = \frac{0.0100 - x}{0.0100 + x} = 0.1778 \Rightarrow x = 0.0041 \text{ mol}$$

$$\text{Moles of HCl added} = 0.0041 \text{ mol} \times 1.00 \text{ L} = 0.0041 \text{ mol}$$

$$\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]} \Rightarrow \frac{[\text{A}^-]}{[\text{HA}]} = 10^{4.0 - 4.75} = 0.1778$$

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