

of substance:

Substance:
 mass of particles of a substance
 in mol

6.02×10^{23} particles
 amount of substance

Constant:
 particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

gas:
 mol, measured in g
 equivalent to M_r

Molar Gas Volume:
 volume of gas per mole of gas molecules in a stated temperature + pressure
 - Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$
 At RTP (20°C and 100kPa):
 (293K and 100000 Pa)
 is $24 \text{ dm}^3 \text{ mol}^{-1}$ at RTP.

Volume /
 dm^3

Electronegativity: → the measure of attraction of a bonded atom in a covalent bond

ELECTRONEGATIVITY INCREASES ACROSS PERIOD:
 - Increase in no. of protons in an atom
 → Nuclear charge increases

ELECTRONEGATIVITY DE:
 - More electron shells
 → More shielding
 → Atomic radius increases
 → Nuclear charge less effective
 → Nuclear attraction less
 → Electronegativity

OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS

Atomic Structure

Electrons:
 → Share p and s but different
 shells different M_r same

Physical Properties:
 → Can be different - eg. different
 melting + boiling points
 → i.e. massy water has higher melting
 point + boiling point + density

Chemical Properties:
 → Same no. of electrons + same electronic
 structure in outer shell + transfer of e^- is
 unaffected hence have same chemical
 properties.
 → Different no. of e^- has no effect
 on bonding + reactions



→ Centre of a small dense
 nucleus of protons + neutrons
 surrounded by electrons orbiting
 in electron orbitals in shells;
 - No. of p^+ = No. of e^-
 → Equal + opposite charges cancel
 so atom has neutral charge
 - Neutrons important in nuclear
 stability - electrostatic repulsion
 between p^+ in small dense nucleus

Mass Number:
 → No. of p^+ + n in nucleus of atom
 - p^+ and n have relative mass of 1
 and e^- mass is negligible to 0

Relative Isotopic Mass:
 → The mass of an atom of an isotope
 relative to $1/12$ th the mass of an atom of $C-12$

Relative Atomic Mass:
 → The weighted mean mass of an atom of
 an element relative to $1/12$ th the mass of
 an atom of $C-12$

Weighted Mean Mass:
 → Average mass of all isotopes, taking into
 account the relative abundance of each isotope

- If dipole
 moment is
 out + molts

e.g. H_2O

- O is more electronegative than H
 - O-H has polar bond so not cancelled out

e.g. CO_2

O is more electronegative than C
 C=O has a dipole moment but cancelled out

Compounds + ions Formulae:

Alkali NO_3^- Halogenoalkane NO_2^-

Ocr Chemistry A Level Unit 2 2014 Paper

Mike Smith



Ocr Chemistry A Level Unit 2 2014 Paper:

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