

# OCR AS Chemistry

## 3.1.2 Group 2

### Topic 23 – Group 2



#### Includes

- Complete lessons
- Animated PowerPoint presentation (46 slides)
- Examples with clear visual explanations
- Worksheet (11 pages)
- Answers to all questions

**Ideal for  
classroom or  
home learning**

#### Animated and interactive slides

##### Reaction of calcium with water

	Ca	+	2 H <sub>2</sub> O	→	Ca(OH) <sub>2</sub>	+	H <sub>2</sub>		oxidised or reduced?
Ca	0				+2				oxidised 0 → +2
H			4 × (+1)		2 × (+1)		2 × 0		reduced +1 → 0
O			2 × (-2)		2 × (-2)				no change
Total	0		0		0		0		redox reaction? yes

**Ca** has been oxidised and **H** has been reduced.  
**O** has not changed oxidation number.

**Solubility of group 2 hydroxides**

Group 2 hydroxides dissolve in water to form alkaline solutions. The more soluble the hydroxide, the more concentrated the OH<sup>-</sup> ions in the resulting solution and the more alkaline it will be (higher pH).

**Trend down group 2 is solubility of hydroxides**

Mg(OH)<sub>2</sub> ↓ solubility increases  
Ca(OH)<sub>2</sub> ↓ solubility increases  
Sr(OH)<sub>2</sub> ↓ solubility increases  
Ba(OH)<sub>2</sub> ↓ solubility increases

**Reaction of calcium oxide with water**

Diagram illustrating the reaction of calcium oxide with water, showing the formation of calcium hydroxide solution and the resulting color change.

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**M Mosston**



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## **Table of Contents Ocr Chemistry A Eee January 20markscheme**

1. Understanding the eBook Ocr Chemistry A Eee January 20markscheme
  - The Rise of Digital Reading Ocr Chemistry A Eee January 20markscheme
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A Eee January 20markscheme
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Chemistry A Eee January 20markscheme
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A Eee January 20markscheme
  - Personalized Recommendations
  - Ocr Chemistry A Eee January 20markscheme User Reviews and Ratings
  - Ocr Chemistry A Eee January 20markscheme and Bestseller Lists
5. Accessing Ocr Chemistry A Eee January 20markscheme Free and Paid eBooks
  - Ocr Chemistry A Eee January 20markscheme Public Domain eBooks
  - Ocr Chemistry A Eee January 20markscheme eBook Subscription Services
  - Ocr Chemistry A Eee January 20markscheme Budget-Friendly Options

6. Navigating Ocr Chemistry A Eee January 20markscheme eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Chemistry A Eee January 20markscheme Compatibility with Devices
  - Ocr Chemistry A Eee January 20markscheme Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Chemistry A Eee January 20markscheme
  - Highlighting and Note-Taking Ocr Chemistry A Eee January 20markscheme
  - Interactive Elements Ocr Chemistry A Eee January 20markscheme
8. Staying Engaged with Ocr Chemistry A Eee January 20markscheme
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry A Eee January 20markscheme
9. Balancing eBooks and Physical Books Ocr Chemistry A Eee January 20markscheme
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry A Eee January 20markscheme
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry A Eee January 20markscheme
  - Setting Reading Goals Ocr Chemistry A Eee January 20markscheme
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry A Eee January 20markscheme
  - Fact-Checking eBook Content of Ocr Chemistry A Eee January 20markscheme
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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