

3. A 0.12 mol dm^{-3} sodium methanoate solution dissociates completely into its ions.

- (a) Formulate the equation, including state symbols, for the equilibrium reaction of the hydrolysis of the methanoate anion. [1]

Sections 1 and 21 of the data booklet may be used for parts (b) to (e).

- (b) Calculate the value of K_a , the acid dissociation constant at 298 K, for an aqueous solution of methanoic acid. [1]

- (c) Calculate the value of K_b , the base dissociation constant, for the conjugate base. [1]

(This question continues on the following page)

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