

New Product Highlights

New VT Range

VT47/570/470/670 Projectors

Stunning new design and high performance at a low price.
Suitable for Education, SME and low-end corporate customers.

The new VT range of LCD projectors breaks new ground with a combination of stunning product design and increased brightness, together with NEC quality and reliability.

The VT Series breaks away from the traditional metal build with a more consumer friendly, light, shiny plastic casing in a sleek, rounded design with soft edges, fitting in with the stylish décor of modern homes and offices. The projectors are easy to set-up and operate, have low fan noise and incorporate advanced features normally found on higher specification models such as Automatic and Manual Vertical Keystone Correction and Wall Colour Correction, and all come with a soft carry case.

VT series are versatile projectors offering high performance at a low cost, and range in brightness and resolution to cover mass market entry level requirements, right up to SME and lower end corporate customers.



VT47

Basic entry-level model for video projection and PowerPoint presentations. Suitable for education and small business markets.

- 1500 ANSI Lumens
- SVGA resolution (800 X 600)
- Short focal lens using digital Zoom
- Direct source keys on top of projector
- RGB output
- Fan noise 33 dB (28 dB in Eco mode)

Special features

- Eco-mode for longer lamp life
- Auto and Manual Keystone Correction
- Off timer function (0.5 to 8 hours)
- 21 on-screen languages
- Manual Wall Colour Correction
- Low fan noise
- Beep sound when operating controls
- Colour management
- 400:1 contrast ratio
- RS232 control

VT570

Entry level model but with high resolution for spreadsheets and more detailed information. Education and small business markets.

- 1300 ANSI Lumens
- XGA resolution (1024 X 768)
- Short focal lens with optical zoom
- 2 x computer/1 video/1 S-Video inputs and 1 RGB output
- Fan noise 33 dB (28 dB in Eco mode)



VT470

With higher brightness and more inputs than the competition, this is a high value SVGA model aimed at SME and low-end corporate customers.

- 2000 ANSI Lumens
- SVGA resolution (800 x 600)
- Short focal lens with optical zoom
- 2 x computer/1 video/1 S-Video inputs and 1 RGB output
- Fan noise 35 dB (30 dB in Eco mode)
- New presentation remote and USB mouse receiver
- Attached 4.5m power cable
- Accessory handle



VT670

High brightness and high resolution with 2 PC inputs and 1 output, makes this a great value mid-range projector for the SME and low-end corporate markets.

- 2000 ANSI Lumens
- XGA resolution (1024 x 768)
- Short focal lens with optical zoom
- 2 x computer/1 video/1 S-Video inputs and 1 RGB output
- Fan noise 35 dB (30 dB in Eco mode)
- New presentation remote and USB mouse receiver
- Attached 4.5m power cable
- Accessory handle



Nec Vt47 Manual

Melchior Lotter



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web google classroom for the following redox equation $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{HSO}_4^{-}$ what will be the stoichiometric coefficient of HSO_4^{-} if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution $\text{Cr}_2\text{O}_7^{2-} + \text{Fe} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{HCN}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{Au} + \text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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